

*Matter and Change, pp. 1-4*

- |    |    |   |
|----|----|---|
| 2. | 5. | c |
| 3. | 5. | c |
| 4. | 7. | c |
| 9. | 3. | a |
10. mass and volume
  11. group or family
  12. homogeneous
  13. products
  14. metal
  15. solid state
  16. chemical properties
  17. atom
  18. compound
  19. nonmetal
  20. mixture
  21. pure substance
  22. basic
  23. metalloid
  24. reactants: carbon and oxygen; product: carbon dioxide
  25. reactant: mercury(II) oxide; products: mercury and oxygen
  26. f
  27. e
  28. c
  29. a
  30. b
  31. i
  32. d
  33. h
  34. e
  35. physical change
  36. chemical change
  37. physical change
  38. physical change
  39. chemical change
  40. physical change
  41. In a solid, particles are packed together in relatively fixed positions. The particles vibrate about a fixed point. In a liquid, particles are close together but can flow around one another. In a gas, particles are at great distances from one another, compared to the particles of liquids and solids.
  42. The composition of a pure substance is the same throughout and does not vary from sample to sample. Pure substances include elements and compounds. A mixture contains more than one substance and can vary in composition from one sample to another and sometimes within different parts of the same sample.
  43. Although energy can be absorbed or released in a

change, it is not destroyed or created. It simply takes a different form.

43. Homogeneous mixtures are uniform in composition; heterogeneous mixtures are not.
44. An element cannot be broken down, but water can be decomposed into hydrogen and oxygen by passing an electrical current through it (electrolysis).

*Measurements and Calculations, pp. 5-8*

- |      |      |
|------|------|
| a    | 5. c |
| 2. c | 6. c |
| 3. c | 7. a |
| 4. a | 8. a |
9. time
  10. mass
  11. density
  12. energy
  13. length or distance
  14. volume
  15. area
  16. qualitative
  17. quantitative
  18. qualitative
  19. quantitative
  20.  $3.00 \times 10^7$  km/s
  21. three
  22. 0.026 g or  $2.6 \times 10^{-2}$  g
  23.  $2.5 \times 10^{-4}$  L
  24. quantity
  25. derived units
  26. conversion factor
  27. 0.0432 kg or  $4.32 \times 10^{-2}$  kg
  28. 5400 mL or  $5.4 \times 10^3$  mL
  29. 300. K
  30.  $1.05 \times 10^3$  J
  31.  $3.51 \times 10^3$  cm
  32. A hypothesis is a testable statement that can be used to make predictions and to carry out further experiments. A theory is a broad generalization that explains a body of facts or phenomena.
  33. Mass is a measure of the quantity of matter. Weight is a measure of the gravitational pull on matter. Mass does not depend on gravitational attraction.
  34. Models are visual, verbal, or mathematical representations. They are used to explain how phenomena occur or how data or events are related.
  35. The graph of two quantities that are directly proportional is a straight line through the origin.

# Chemistry A Study Of Matter 114 Answer Key

**K Morrison**



**Chemistry A Study Of Matter 114 Answer Key:**

## **Chemistry A Study Of Matter 114 Answer Key** Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has become much more apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Chemistry A Study Of Matter 114 Answer Key**," compiled by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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