

Stoichiometry Worksheet # 1

1) Glucose is used as a source of energy by the human body. The overall reaction in the body is



Calculate the number of moles of oxygen needed to oxidize 12.5 g of $\text{C}_6\text{H}_{12}\text{O}_6$.

2) Ammonia is synthesized from hydrogen and nitrogen according to the following equation.



If an excess of nitrogen is reacted with 3.41 moles of hydrogen gas, how many grams of ammonia (NH_3) can be produced?

3) Assume that in the decomposition of potassium chlorate, KClO_3 ,



80.5 moles of O_2 form. How many grams of potassium chlorite, the other product would be formed?

4) In a single displacement reaction,



9.23 moles of aluminum react with excess hydrochloric acid. How many moles of hydrogen will be produced?

5) The compound "Cisplatin" $\text{PtCl}_2(\text{NH}_3)_2$, has been found to be effective in treating some types of cancer. It can be synthesized using the following reaction



A. How many moles of "cisplatin" can be produced from 2.50 g K_2PtCl_4 ?

B. How many moles of NH_3 would be needed if 3.8 moles of K_2PtCl_4 were present?

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Wenbin Ji



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Table of Contents Chemistry Chapter 12 Study For Content Mastery Stoichiometry Answers

1. Understanding the eBook Chemistry Chapter 12 Study For Content Mastery Stoichiometry Answers
 - The Rise of Digital Reading Chemistry Chapter 12 Study For Content Mastery Stoichiometry Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chapter 12 Study For Content Mastery Stoichiometry Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Chapter 12 Study For Content Mastery Stoichiometry Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Chapter 12 Study For Content Mastery Stoichiometry Answers
 - Personalized Recommendations

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- Chemistry Chapter 12 Study For Content Mastery Stoichiometry Answers User Reviews and Ratings
- Chemistry Chapter 12 Study For Content Mastery Stoichiometry Answers and Bestseller Lists
- 5. Accessing Chemistry Chapter 12 Study For Content Mastery Stoichiometry Answers Free and Paid eBooks
 - Chemistry Chapter 12 Study For Content Mastery Stoichiometry Answers Public Domain eBooks
 - Chemistry Chapter 12 Study For Content Mastery Stoichiometry Answers eBook Subscription Services
 - Chemistry Chapter 12 Study For Content Mastery Stoichiometry Answers Budget-Friendly Options
- 6. Navigating Chemistry Chapter 12 Study For Content Mastery Stoichiometry Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Chapter 12 Study For Content Mastery Stoichiometry Answers Compatibility with Devices
 - Chemistry Chapter 12 Study For Content Mastery Stoichiometry Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Chapter 12 Study For Content Mastery Stoichiometry Answers
 - Highlighting and Note-Taking Chemistry Chapter 12 Study For Content Mastery Stoichiometry Answers
 - Interactive Elements Chemistry Chapter 12 Study For Content Mastery Stoichiometry Answers
- 8. Staying Engaged with Chemistry Chapter 12 Study For Content Mastery Stoichiometry Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Chapter 12 Study For Content Mastery Stoichiometry Answers
- 9. Balancing eBooks and Physical Books Chemistry Chapter 12 Study For Content Mastery Stoichiometry Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Chapter 12 Study For Content Mastery Stoichiometry Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Chapter 12 Study For Content Mastery Stoichiometry Answers
 - Setting Reading Goals Chemistry Chapter 12 Study For Content Mastery Stoichiometry Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Chapter 12 Study For Content Mastery Stoichiometry Answers
 - Fact-Checking eBook Content of Chemistry Chapter 12 Study For Content Mastery Stoichiometry Answers

- Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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