

Name: _____

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Chapter 6 Study Guide: Periodicity

1. Describe the historical development of the periodic table (Newland, Meyer, Mendeleev, Mosely) Describe the contributions made by each of the scientist.

Newland – First to order the elements in a periodic table –developed the law of octaves

Meyers – Ordered elements by atomic mass but is given little credit for the current periodic table

Mendeleev – ordered elements by atomic mass and is given the most credit for the current periodic table because he left holes in his table for undiscovered elements and described their properties

Mosely – ordered elements by atomic number (modern periodic table)

2. Fill in the blanks below:

The elements in the periodic table are arranged in order of increasing Atomic

Number. The vertical columns are called Groups or

Families. The horizontal rows are called Periods. The

elements within the same Family have the same chemical properties due to the number of their Valence Electrons. Metals are located on the

Left side of the periodic table. The nonmetals are located on the

Right side of the periodic table. The metalloids are located on

the Stair step between the metals and nonmetals.

3. Describe the general properties of the following groups of elements on the periodic table:

- a. Alkali Metals

1 valence electrons, metals, highly reactive, s-block

- b. Alkaline Earth metals

2 valence electrons, metals, reactive, s-block

- c. Halogens

7 valence electrons, high reactive, nonmetals, p-block

- d. Noble Gases

8 noble electrons, nonreactive, nonmetals, all gases, p-block

- e. Transition Metals

Variable valence electrons, metals, malleable, ductile, metals, conductive

- f. Inner Transition Metals

Variable valence electrons, metals, f-block

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