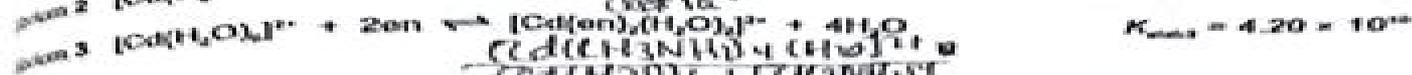
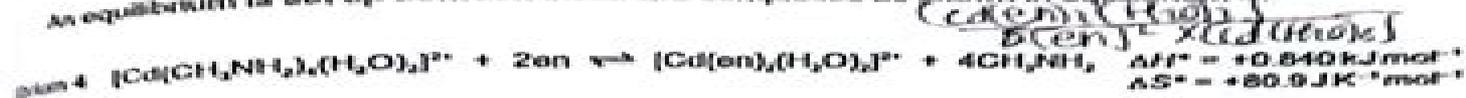


(b) Cadmium ions form complexes with methylamine,  $\text{CH}_3\text{NH}_2$ , and with 1,2-diaminoethane,  $\text{H}_2\text{NCH}_2\text{CH}_2\text{NH}_2$ , as shown in equilibria 2 and 3. 1,2-diaminoethane is shown as en.



An equilibrium is set up between these two complexes as shown in equilibrium 4.



(i)  $K_{\text{eq}}$  is the equilibrium constant for equilibrium 4.

Write an expression for  $K_{\text{eq}}$  in terms of  $K_{\text{stab}2}$  and  $K_{\text{stab}3}$ .

$$K_{\text{eq}} = \frac{[\text{CH}_3\text{NH}_2]^4 \times [\text{Cd}(\text{en})_2(\text{H}_2\text{O})_2]^{2+}}{[\text{en}]^2 \times [\text{Cd}(\text{CH}_3\text{NH}_2)_4(\text{H}_2\text{O})_2]^{2+}} = \frac{K_{\text{stab}3}}{K_{\text{stab}2}}$$

(ii) Calculate the value of the standard Gibbs free energy change,  $\Delta G^\circ$ , for equilibrium 4 at 298 K.

$$\Delta G = \Delta H - T\Delta S = +0.840 - 298 \times (80.9 \times 10^{-3})$$

$$\Delta G^\circ = -23.3 \text{ kJ mol}^{-1} \quad [2]$$

(iii) State how the value of  $\Delta G^\circ$  changes as the temperature increases. Explain your answer.

As temp increases,  $\Delta G$  value gets more negative as value for  $(T\Delta S)$  in the eqn increases. [1]

[Total: 9]

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