

# Unit 1 – Atomic Structure MC Regents Questions Key

3

1. Which element has an atom with the electron configuration 2-8-8-2?

- |       |       |
|-------|-------|
| 1. Mg | 3. Ca |
| 2. Ni | 4. Ge |

**3** An atom of the element has 20 electrons (2 + 8 + 8 + 2) and 20 protons. Use the Periodic Table of the Elements and find the electron configuration of the element whose atomic number is 20.

1

2. Which type of radiation has *neither* mass nor charge?

- |            |          |
|------------|----------|
| 1. gamma   | 3. alpha |
| 2. neutron | 4. beta  |

**1** All forms of electromagnetic radiation, such as gamma radiation, have neither mass nor electric charge.

3

3. Which list of particles is in order of increasing mass?

- (1) proton → electron → alpha particle
- (2) proton → alpha particle → electron
- (3) electron → proton → alpha particle
- (4) alpha particle → electron → proton

**3** An electron is the least massive of the three particles; a proton is approximately 2,000 times more massive than an electron; an alpha particle is the nucleus of a helium-4 atom and is approximately 4 times more massive than a proton.

3

4. Compared to a sodium atom in the ground state, a sodium atom in the excited state must have

- |                                  |                                    |
|----------------------------------|------------------------------------|
| 1. a greater number of electrons | 3. an electron with greater energy |
| 2. a smaller number of electrons | 4. an electron with less energy    |

**3** Atoms become excited by absorbing energy. Therefore, an excited atom has more energy than it had in the ground state.

**Wrong Choices Explained:**

(1), (2) An excited atom is still a neutral atom. It has neither greater nor fewer electrons than an atom in the ground state.

1

5. Which particles are isotopes of each other?

# Chemistry Regents Questions And Answers Atomic Structure

**Katrin Zwirglmaier**



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### **Table of Contents Chemistry Regents Questions And Answers Atomic Structure**

1. Understanding the eBook Chemistry Regents Questions And Answers Atomic Structure
  - The Rise of Digital Reading Chemistry Regents Questions And Answers Atomic Structure
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Regents Questions And Answers Atomic Structure
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an eBook Chemistry Regents Questions And Answers Atomic Structure
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Regents Questions And Answers Atomic Structure
  - Personalized Recommendations

- Chemistry Regents Questions And Answers Atomic Structure User Reviews and Ratings
- Chemistry Regents Questions And Answers Atomic Structure and Bestseller Lists
- 5. Accessing Chemistry Regents Questions And Answers Atomic Structure Free and Paid eBooks
  - Chemistry Regents Questions And Answers Atomic Structure Public Domain eBooks
  - Chemistry Regents Questions And Answers Atomic Structure eBook Subscription Services
  - Chemistry Regents Questions And Answers Atomic Structure Budget-Friendly Options
- 6. Navigating Chemistry Regents Questions And Answers Atomic Structure eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry Regents Questions And Answers Atomic Structure Compatibility with Devices
  - Chemistry Regents Questions And Answers Atomic Structure Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Regents Questions And Answers Atomic Structure
  - Highlighting and Note-Taking Chemistry Regents Questions And Answers Atomic Structure
  - Interactive Elements Chemistry Regents Questions And Answers Atomic Structure
- 8. Staying Engaged with Chemistry Regents Questions And Answers Atomic Structure
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Regents Questions And Answers Atomic Structure
- 9. Balancing eBooks and Physical Books Chemistry Regents Questions And Answers Atomic Structure
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Regents Questions And Answers Atomic Structure
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Regents Questions And Answers Atomic Structure
  - Setting Reading Goals Chemistry Regents Questions And Answers Atomic Structure
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Regents Questions And Answers Atomic Structure
  - Fact-Checking eBook Content of Chemistry Regents Questions And Answers Atomic Structure

- Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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