

Introduction 11

An Electron's Address

Name: _____
 Date: _____
 Class: _____

Introduction: Energy Levels and Electrons

The electrons in the following model would have populated the following lowest energy levels. The electrons would fill the lowest one's first, then the next one.

Remember that lowest energy levels are the most fundamental lowest energy levels. Each complete level is called a shell and containing orbitals.

Figure 1

| Energy Level | Number of electrons in the lowest energy level |
|--------------|--|
| 1 | 2 |
| 2 | 8 |
| 3 | 18 |
| 4 | 32 |

Note that there is a specific thing we call a "shell" or "level" of the n^{th} energy level because there are only a few orbitals in each of the n^{th} energy level.

Critical Thinking Questions

- How many electrons exist in the 1^{st} energy level?
 The number of orbitals in the 1^{st} energy level is one which is the only one orbital.
- How many electrons exist in the 2^{nd} energy level?
 The number of orbitals in the 2^{nd} energy level is eight which is the only one orbital.
- How many electrons exist in the 3^{rd} energy level?
 The number of orbitals in the 3^{rd} energy level is 18 which is the only one orbital.
- How many electrons would you expect to have in the 4^{th} energy level?
 The number of orbitals in the 4^{th} energy level would be 32 orbitals. There would be 18 orbitals which is the number of orbitals in the energy level below.
- How do 18 orbitals exist in the 3^{rd} energy level?
 The 18 orbitals exist in the 3^{rd} energy level because there are 18 orbitals in the 3^{rd} energy level. The other orbitals exist in the 1^{st} energy level (the 1st shell) and the

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