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# AP Biology Unit 2 Test Questions with 100% Correct Answers

2 organelles that contain their own DNA separate from the chromosome Correct Answer Mitochondria and chloroplast

What domains fall under prokaryotic cells? Correct Answer Bacteria and archaea

What are the basic features of all cells? Correct Answer Membrane, cytoplasm, chromosomes, ribosomes

This organelle synthesizes lipids, metabolizes carbohydrates, detoxifies poison, and stores calcium Correct Answer Smooth ER

This cell is characterized by having DNA in a nucleus that is bound by a nuclear envelope and membrane bound organelles Correct Answer Eukaryotic cell

This is a membranous sac of hydrolytic enzymes that can digest macromolecules Correct Answer Lysosome

The \_\_\_\_\_ is a selective barrier that allows oxygen, nutrients, and wastes to service the volume of every cell Correct Answer Plasma membrane

Ribosomes use the information from the DNA to make what Correct Answer Proteins

This cell type is characterized by having no nucleus, but instead a nucleoid, and no membrane bound organelles Correct Answer Prokaryotic cell

Diffusion through protein channels is known as what? Correct Answer Facilitated diffusion

Site of ribosomal RNA synthesis Correct Answer Nucleolus

What happens to a plant cell when placed in pure water (hypotonic solution)? Correct Answer It becomes turgid, cell wall protects from bursting

Site of photosynthesis Correct Answer chloroplast

These vacuoles pump excess water out of the cell Correct Answer Contractile

This ER lacks ribosomes Correct Answer Smooth ER

This organelle modifies products of the ER, manufactures certain macromolecules, sorts and packages materials into transport vesicles Correct Answer Golgi apparatus

# Dna Biology Test With Answers

**SA Dillow**



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web google classroom for the following redox equation  $\text{MnO}_4^- + \text{SO}_2 + 2\text{H}^+ \rightarrow \text{Mn}^{2+} + \text{HSO}_4^-$  what will be the stoichiometric coefficient of  $\text{HSO}_4^-$  if the above equation is balanced in an acidic medium stuck use a hint report a problem do 4 problems

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web may 25 2021 balancing redox equations using half reactions another way to balance redox reactions is by the half reaction method this technique involves breaking an equation into its two separate components the oxidation reaction and

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web balance this redox reaction by the half reaction method by assuming an acidic solution  $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} \rightarrow \text{Cr}^{3+} + \text{Fe}^{3+}$  balance the redox reaction in exercise 9 by the half reaction method by assuming a basic solution

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web solution 1 the two half reactions balanced as if in acidic solution  $2\text{CN}^- + \text{Au} + \text{CN}^- + 2\text{e}^- + 2\text{H}^+ + \text{O}_2 \rightarrow \text{H}_2\text{O}_2 + 2\text{CNO}^-$  make electrons equal convert to basic solution  $4\text{CN}^- + 2\text{Au} + 2\text{CN}^- + 2\text{e}^-$  multiplied by a factor of 2  $2\text{e}^- + 2\text{H}_2\text{O} + \text{O}_2 \rightarrow \text{H}_2\text{O}_2 + 2\text{OH}^-$  the final answer

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web aug 29 2023 the first step to balance any redox reaction is to separate the reaction into half reactions the substance being reduced will have electrons as reactants and the oxidized substance will have electrons as products usually all reactions are written as reduction reactions in half reaction tables

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web apr 17 2013 in order to balance redox reactions well you must first be able to assign oxidation numbers well oxidation the loss of electrons resulting in a more positively charged species reduction the gain of electrons resulting in a

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web a redox equation can be balanced using the following stepwise procedure 1 divide the equation into two half reactions 2 balance each half reaction for mass and charge 3 equalize the number of electrons transferred in each half reaction 4

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web the half equation method separates the oxidation and reduction of a redox reaction in half reactions overall scheme for the half reaction method step 1 split reaction into half reactions reduction and oxidation step 2 balance the charge or oxidation number with electrons step 3 balance ce o by adding ce h<sub>2</sub>o

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web when balancing equations for redox reactions occurring in basic solution it is often necessary to add oh ions or the oh h<sub>2</sub>o pair to fully balance the equation in this video we ll walk through this process for the reaction between clo and cr oh<sub>4</sub> in basic solution created by jay

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web bonus problem here is a balanced net ionic equation  $3\text{Br}_2 + 6\text{OH}^- \rightarrow 5\text{Br}^- + \text{BrO}_3^- + 3\text{H}_2\text{O}$  the above equation takes place in two stages first bromine reacts with oh ions to form bromide ions and bro ions second the bro ions react to form bromide ions and bromate v ions

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web mar 13 2023 balancing redox equations by the ion electron method there are two principal methods for balancing redox equations oxidation state method ion electron method the latter is easier to use with redox reactions in aqueous solution and if necessary can be adapted to many situations that are not in aqueous solution

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation  $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$  now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balancing redox reactions in acidic solution problems 1 10 problem 1  $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} + \text{Cr}^{3+} + \text{Fe}^{3+}$  solution 1 balanced half reactions  $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$   $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$  2 equalize the electrons  $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$   $6\text{Fe}^{2+} \rightarrow 6\text{Fe}^{3+} + 6\text{e}^-$  multiplied by a factor of 6

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web when balancing equations for redox reactions occurring in acidic solution it is often necessary to add h ions or the h h<sub>2</sub>o pair to fully balance the equation in this video we ll walk through this process for the reaction between dichromate cr<sub>2</sub>o<sub>7</sub><sup>2-</sup> and chloride cl ions in acidic solution created by jay

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web balance each half reaction in basic solution  $4\text{S} + 6\text{H}_2\text{O} + 7\text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{NO}_3^- + \text{SO}_4^{2-} + 2\text{MnO}_2 + \text{Cr}^{3+} + \text{NO}_3^- + \text{SO}_2 + \text{Mn}_2\text{O}_3$  balance each redox

reaction in acid solution using the half reaction method  
8  $\text{H}_2\text{O}_2$   $\text{Cr}_2\text{O}_7^{2-}$  9  $\text{TeO}_3$  2  $\text{N}_2\text{O}_4$   $\text{O}_2$   $\text{Cr}^{3+}$  10  $\text{ReO}_4^-$  11  $\text{PbO}_2$   $\text{IO}_2^-$   $\text{Pb}^{2+}$  12  
as  $\text{NO}_3^-$   $\text{IO}_3^-$   $\text{Re}^{+3}$   $\text{H}_2\text{AsO}_4^-$

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web aug 29 2023 some points to remember when balancing redox reactions the equation is separated into two half equations one for oxidation and one for reduction the equation is balanced by adjusting coefficients and adding  $\text{H}^+$   $\text{OH}^-$  and  $\text{e}^-$  in this order balance the atoms in the equation apart from  $\text{O}$  and  $\text{H}$

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively  $\text{C}_6\text{H}_8\text{O}_6$   $2\text{HNO}_2$   $\text{C}_6\text{H}_6\text{O}_6$   $2\text{H}_2\text{O}_2$  choose 1 answer  $\text{C}_6\text{H}_8\text{O}_6$  and  $\text{NO}_2^-$   $\text{C}_6\text{H}_8\text{O}_6$  and  $\text{NO}_2$   $\text{NO}_2$  and  $\text{C}_6\text{H}_8\text{O}_6$   $\text{NO}_2^-$  and  $\text{C}_6\text{H}_8\text{O}_6$   $\text{H}^+$  and  $\text{NO}_2^-$   $\text{C}_6\text{H}_8\text{O}_6$  and  $\text{NO}_2$   $\text{C}_6\text{H}_8\text{O}_6$  and  $\text{NO}_2^-$   $\text{H}^+$  and  $\text{NO}_2^-$