

## Chapter 4 – Atomic Structure

**Short Answer:** Answer the following questions in complete sentences using science vocabulary. Make sure that your answers are complete and detailed.

1. A sample of calcium contains calcium-40, calcium-44, calcium-42, calcium-48, calcium-43, and calcium-46 atoms. Explain why these atoms have different mass numbers but must have the same atomic number.

*The atoms have different mass numbers because they have a different number of neutrons. The mass number is equaled to the number of protons plus neutrons.*

*The atoms must have the same atomic number because the atomic number is equaled to the number of protons. The number of protons is unique for each element.*

2. Complete the following table. Circle the name of the element(s) that are isotopes. Complete Bohr's Model diagrams for the element with eleven electrons and the element with an atomic number of 9.

Element Name	Element Symbol	Atomic Number	Mass Number	Protons	Neutrons	Electrons
Scandium	Sc	21	45	21	24	21
Fluorine	F	9	19	9	10	9
Sodium	Na	11	24	11	13	11
Answers Vary					82	

# Chapter 4 Atomic Structure Practice Problems Answers

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Rogers**



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Survival Guide to General Chemistry Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics. The focus is not to recap or review the theoretical concepts well described in the available texts. Instead, the topics and descriptions in this book make available specific, detailed, step-by-step methods and procedures for solving the major types of problems in general chemistry. Explanations, instructional process sequences, solved examples, and completely solved practice problems are greatly expanded, containing significantly more detail than can usually be devoted to in a comprehensive text. Many chapters also provide alternative viewpoints as an aid to understanding. Key Features: The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester. Each is written in a specific and detailed step-by-step process for problem solving, whether mathematical or conceptual. Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts. Includes a chapter designed to eliminate confusion concerning acid-base reactions, which often persists through working with acid-base equilibrium. Many chapters provide alternative viewpoints as an aid to understanding. This book addresses a very real need for a large number of incoming freshmen in STEM fields. *MCAT General Chemistry Review, 3rd Edition* The Princeton Review, 2016-01-05. Make sure you're studying with the most up-to-date prep materials. Look for the newest edition of this title. The Princeton Review *MCAT General Chemistry Review 4th Edition* ISBN 9780593516256 on sale November 2022. Publisher's Note: Products purchased from third-party sellers are not guaranteed by the publisher for quality or authenticity and may not include access to online tests or materials included with the original product. *Chemistry 'O' Level* Rex M. Heyworth, 2007. *Student's Guide to Chemistry; a Modern Introduction* David Brooks, 1974. *Foundations of College Chemistry, Alternate* Morris Hein, Susan Arena, 2010-01-26. Learning the fundamentals of chemistry can be a difficult task to undertake for health professionals. For over 35 years, this book has helped them master the chemistry skills they need to succeed. It provides them with clear and logical explanations of chemical concepts and problem solving. They'll learn how to apply concepts with the help of worked-out examples. In addition, *Chemistry in Action* features and conceptual questions checks brings together the understanding of chemistry and relates chemistry to things health professionals experience on a regular basis. **Chemistry is Phenomenal** Tracy Poulsen, **Cracking the GRE Chemistry Test**, 2002. The Princeton Review realizes that acing the GRE Chemistry Test is very different from getting straight A's in school. We don't try to teach you everything there is to know about chemistry; only the techniques you'll need to score higher on the test. There's a big difference. In *Cracking the GRE Chemistry Test*, we'll teach you how to think like the test writers and eliminate answer choices that look right but are planted to fool you. Raise your score by focusing on the material most likely to appear on the test. Test your knowledge with review questions for each chemistry topic covered. This book includes one full-length practice GRE Chemistry Test. All of our practice

questions are like the ones you'll see on the actual GRE Chemistry Test and we fully explain every solution

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Merrill Chemistry Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998

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**Study Guide** Darrell D. Ebbing, Susan M. Schelble, 1998

*Student's Guide to Introduction to Chemical Principles* by Edward I. Peters, 2d Ed Peter P. Berlow, 1978

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**High School Chemistry Teachers Magazine**, 1967

*McGraw-Hill's ASVAB Basic Training for the AFQT* Janet E. Wall, 2006-01-17 Inside information on the verbal and math skill levels required for specific military occupations Author was formerly a key member of the ASVAB team at the Department of Defense Ready for publication just when the armed forces are raising their enlistment goals

**Problem Solving Guide and Workbook for Introductory Chemistry** by Steve Russo, Mike Silver Sandra Yancy McGuire, Steve

Russo, Mike Silver, 2002 Provides over 175 worked examples and more than 500 practice problems and quiz questions to help students develop and practice their problem solving skills Illinois Technograph, 1961 Fundamentals of Chemistry  
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