

Chapter 4 – Atomic Structure

Short Answer: Answer the following questions in complete sentences using science vocabulary. Make sure that your answers are complete and detailed.

1. A sample of calcium contains calcium-40, calcium-44, calcium-42, calcium-48, calcium-43, and calcium-46 atoms. Explain why these atoms have different mass numbers but must have the same atomic number.

The atoms have different mass numbers because they have a different number of neutrons. The mass number is equaled to the number of protons plus neutrons.

The atoms must have the same atomic number because the atomic number is equaled to the number of protons. The number of protons is unique for each element.

2. Complete the following table. Circle the name of the element(s) that are isotopes. Complete Bohr's Model diagrams for the element with eleven electrons and the element with an atomic number of 9.

| Element Name | Element Symbol | Atomic Number | Mass Number | Protons | Neutrons | Electrons |
|--------------|----------------|---------------|-------------|---------|----------|-----------|
| Scandium | Sc | 21 | 45 | 21 | 24 | 21 |
| Fluorine | F | 9 | 19 | 9 | 10 | 9 |
| Sodium | Na | 11 | 24 | 11 | 13 | 11 |
| Answers Vary | | | | | 82 | |

Chapter 4 Atomic Structure Test Answer Key

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