

Chapter 6 Covalent Bonds

Section 6.2 Covalent Bonding**(pages 166–188)**

This section discusses the formation of covalent bonds and the factors that determine whether a molecule is polar or nonpolar. It also discusses attractions between molecules.

Reading Strategy (page 166)

Relating Text and Visuals As you read the section, look closely at Figure 9. Complete the table by describing each type of model shown. For more information on this Reading Strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.

Molecular Models	
Model	Description
Electron-dot diagrams	Dots represent valence electrons.
Structural formula	
Space-filling	
Electron cloud	

Covalent Bonds (pages 166–187)

- Define a covalent bond. _____

- A molecule is a _____ group of atoms that are joined together by one or more covalent bonds. Circle the correct answer.
negative neutral positive
- Is the following sentence true or false? In a covalent bond, the atoms are held together by the attractions between the shared electrons and the protons in each nucleus. _____
- Circle the correct answer. Nitrogen has five valence electrons. How many pairs of electrons must two nitrogen atoms share in order for each atom to have eight valence electrons?
 - one
 - two
 - three

Chapter 6 Chemical Bonding Section 2 Covalent Answer Key

ML Yell



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Focus on Physical Science California Edition Michael J. Padilla, 2007 **Physical Biochemistry** David Sheehan, 2013-04-30 As will be seen there is not much missing here I thought that the sections were well balanced with rarely too much or too little on a given topic This is a text to be welcomed by both teachers and students BIOCHEMISTRY MOLECULAR BIOLOGY EDUCATION on the first edition The second edition of this successful textbook explains the basic principles behind the key techniques currently used in the modern biochemical laboratory and describes the pros and cons of each technique and compares one to another It is non mathematical comprehensive and approachable for students who are not physical chemists A major update of this comprehensive accessible introduction to physical biochemistry Includes two new chapters on proteomics and bioinformatics Introduces experimental approaches with a minimum of mathematics and numerous practical examples Provides a bibliography at the end of each chapter Written by an author with many years teaching and research experience this text is a must have for students of biochemistry biophysics molecular and life sciences and food science Addison-Wesley Introduction to Physical Science Michael B. Leyden, 1988 **Chemical Interactions**, 2005 *General Chemistry* Kenneth W. Whitten, Raymond E. Davis, 1996 Are you looking for the key to success in your chemistry class In CHEMISTRY you will find a strong molecular reasoning focus problem solving exercises and an innovative online homework management system that will prepare you for any challenge you might encounter The textbook is filled with learning aids that will help you master concepts of the course **Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science**, 2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction

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