

CHAPTER 2: ATOMS AND THE PERIODIC TABLE

1) 2.1 Elements

- a) **Element**: pure substance that cannot be broken down into simpler substances by a chemical reaction
- i) Identified by a one- or two-letter symbol
- b) **Periodic table**: a schematic way of organizing groups of elements that have similar chemical properties

LEFT SIDE

Metals, Nonmetals, and Metalloids

The periodic table is color-coded as follows:

- Metals (pink):** Includes groups 1-10, 11-12, and 13-16.
- Nonmetals (yellow):** Includes groups 17 and 18.
- Metalloids (light blue):** Includes groups 13-16.

Solid @ room temp. (besides Hg)

- i) **Metals**: conduct heat & electricity. Ductile, meaning they can be drawn into wires, and malleable, meaning they can be hammered into shapes

- ii) **Metalloids**: have properties intermediate between metals and nonmetals

- iii) **Nonmetals**: poor conductors of heat and electricity *may be solid, liquid, or gas*

RIGHT SIDE

- c) **Compound**: pure substance formed by chemically combining two or more elements together

- i) **Chemical formula**: uses element symbols to show the identity of the elements forming a compound & subscripts to show the ratio of atoms in the compound



(1) Molecular art can be used to illustrate the composition & state of elements & compounds

(a) Black: carbon ; white: hydrogen ; red: oxygen ; blue: nitrogen

2) 2.2 Structure of the Atom

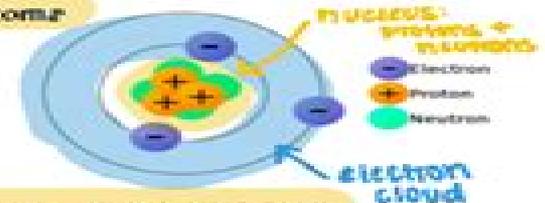
- a) All matter is composed of the same basic building blocks called atoms

- i) **Atom** composed of three subatomic particles:

- (1) **Proton** has a positive (+) charge
- (2) **Electron** has a negative (-) charge
- (3) **Neutron** has no charge

~1 amu

- ii) Protons & neutrons have approximately the same mass, electron mass is much less



ELEMENTS OF LIFE:
 Carbon, Oxygen, Hydrogen, + Nitrogen
 → makes up 98% of the mass of the human body
 → building blocks

Chapter Test Atom And Periodic Table

G Thomas



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