

QUALITATIVE ANALYSIS

Experiment	Observation	Inference
<u>Preliminary Test</u>		
1. Noted the colour of the given salt	<ol style="list-style-type: none"> 1. Blue or Bluish Green 2. Greenish 3. Light Green 4. Dark Brown 5. Light Pink 6. Colourless 	<p>Cu^{2+} or Ni^{2+} Ni^{2+} Fe^{2+} Fe^{3+} Mn^{2+} Absence of Cu^{+2}, Ni^{2+}, Fe^{2+}, Fe^{3+}, Mn^{2+}</p>
2. Noted the solubility	<ol style="list-style-type: none"> 1. Soluble in water 2. Soluble in dilute HCl 	<p>NH_4^+ may be present Absence of group I cations</p>
3. Dry Heating Test: Heated a pinch of the salt in a dry Test tube	<ol style="list-style-type: none"> 1. Colour, odourless gas 2. Colourless gas with rotten egg smell 3. Colourless gas with vinegar smell 4. Colourless gas with ammonia smell 5. Reddish brown gas 6. Greenish yellow gas 7. Violet gas 	<p>CO_3^{2-} may be present S^{2-} may be present CH_3COO^- may be present NH_4^+ may be present NO_3^- or Br^- may be present Cl^- may be present I^- may be present</p>
<u>Identification of acid radical (Anion)</u>		
1. Dilute Acid Test: Treated a pinch of the substance With dilute H_2SO_4	<ol style="list-style-type: none"> 1. Colourless, odourless gas with brisk effervescence which turned lime water milky. 2. Vinegar Smell 3. No characteristic Reaction 	<p>CO_3^{2-} present $\text{CO}_3^{2-} + 2\text{H}^+ \rightarrow \text{CO}_2 + \text{H}_2\text{O}$ $\text{Ca}(\text{OH})_2 + \text{CO}_2 \rightarrow \text{CaCO}_3 + \text{H}_2\text{O}$</p> <p>$\text{CH}_3\text{COO}^-$ Present $\text{CH}_3\text{COO}^- + \text{H}^+ \rightarrow \text{CH}_3\text{COOH}$</p> <p>Absence of CO_3^{2-}, CH_3COO^-</p>
2. Concentrated Acid Test: Treated a pinch of the substance With concentrated H_2SO_4	<ol style="list-style-type: none"> 1. Colourless gas with pungent Smell which gave dense white fumes when a glass rod dipped in NH_4OH was brought near the mouth of the test tube 2. Reddish brown vapours with Pungent smell. Solution turns red 3. Deep violet vapours with pungent smell. Sides of test tube 	<p>Cl^- present $\text{Cl}^- + \text{H}^+ \rightarrow \text{HCl}$ $\text{HCl} + \text{NH}_4\text{OH} \rightarrow \text{NH}_4\text{Cl} + \text{H}_2\text{O}$</p> <p>$\text{Br}^-$ present $\text{Br}^- + \text{H}^+ \rightarrow \text{HBr}$ $2\text{HBr} + \text{H}_2\text{SO}_4 \rightarrow 2\text{H}_2\text{O} + \text{Br}_2 + \text{SO}_2$</p> <p>$\text{I}^-$ present $\text{I}^- + \text{H}^+ \rightarrow \text{HI}$ $2\text{HI} + \text{H}_2\text{SO}_4 \rightarrow 2\text{H}_2\text{O} + \text{I}_2 + \text{SO}_2$</p>

Chemistry Practical Qualitative Analysis Sheet

Terry C. Jones



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Table of Contents Chemistry Practical Qualitative Analysis Sheet

1. Understanding the eBook Chemistry Practical Qualitative Analysis Sheet
 - The Rise of Digital Reading Chemistry Practical Qualitative Analysis Sheet
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Practical Qualitative Analysis Sheet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Practical Qualitative Analysis Sheet
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Practical Qualitative Analysis Sheet
 - Personalized Recommendations
 - Chemistry Practical Qualitative Analysis Sheet User Reviews and Ratings
 - Chemistry Practical Qualitative Analysis Sheet and Bestseller Lists
5. Accessing Chemistry Practical Qualitative Analysis Sheet Free and Paid eBooks
 - Chemistry Practical Qualitative Analysis Sheet Public Domain eBooks
 - Chemistry Practical Qualitative Analysis Sheet eBook Subscription Services
 - Chemistry Practical Qualitative Analysis Sheet Budget-Friendly Options

6. Navigating Chemistry Practical Qualitative Analysis Sheet eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Practical Qualitative Analysis Sheet Compatibility with Devices
 - Chemistry Practical Qualitative Analysis Sheet Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Practical Qualitative Analysis Sheet
 - Highlighting and Note-Taking Chemistry Practical Qualitative Analysis Sheet
 - Interactive Elements Chemistry Practical Qualitative Analysis Sheet
8. Staying Engaged with Chemistry Practical Qualitative Analysis Sheet
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Practical Qualitative Analysis Sheet
9. Balancing eBooks and Physical Books Chemistry Practical Qualitative Analysis Sheet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Practical Qualitative Analysis Sheet
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Practical Qualitative Analysis Sheet
 - Setting Reading Goals Chemistry Practical Qualitative Analysis Sheet
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Practical Qualitative Analysis Sheet
 - Fact-Checking eBook Content of Chemistry Practical Qualitative Analysis Sheet
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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