

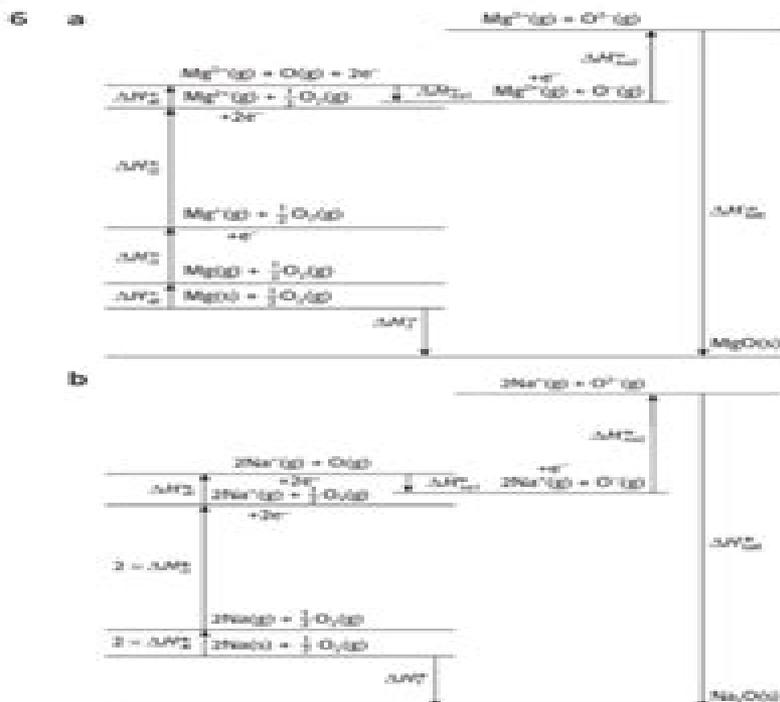
Answers to SAQs

Chapter 19

- 1 a standard temperature = 298 K, standard pressure = 1×10^5 Pa / 100 kPa
- b i $\text{Mg}^{2+}(\text{g}) + \text{O}^{2-}(\text{g}) \rightarrow \text{MgO}(\text{s})$
 ii $\text{K}^+(\text{g}) + \text{Br}^-(\text{g}) \rightarrow \text{KBr}(\text{s})$
 iii $2\text{Na}^+(\text{g}) + \text{S}^{2-}(\text{g}) \rightarrow \text{Na}_2\text{S}(\text{s})$
- 2 a The bond energy for chlorine is the enthalpy change $\text{Cl}_2(\text{g}) \rightarrow 2\text{Cl}(\text{g})$. The enthalpy change of atomisation is $\frac{1}{2}\text{Cl}_2(\text{g}) \rightarrow \text{Cl}(\text{g})$. So the enthalpy change of atomisation is $+244/2 = +122 \text{ kJ mol}^{-1}$.
- b i $\frac{1}{2}\text{O}_2(\text{g}) \rightarrow \text{O}(\text{g})$
 ii $\text{Ba}(\text{s}) \rightarrow \text{Ba}(\text{g})$
 iii $\frac{1}{2}\text{Br}_2(\text{l}) \rightarrow \text{Br}(\text{g})$
- c 0 kJ mol^{-1}
 Because helium exists naturally as single gaseous atoms, no change is involved in the process $\text{He}(\text{g}) \rightarrow \text{He}(\text{g})$.
- 3 a There must be an input of energy to overcome the repulsive forces between the (negative) electron and the negative ion.
- b $+440 \text{ kJ mol}^{-1}$
- c i $\text{I}(\text{g}) + \text{e}^- \rightarrow \text{I}^-(\text{g})$
 ii $\text{S}^-(\text{g}) + \text{e}^- \rightarrow \text{S}^{2-}(\text{g})$
- 4 a i $\text{Cs}(\text{g}) \rightarrow \text{Cs}^+(\text{g}) + \text{e}^-$
 ii $\text{Al}^{3+}(\text{g}) \rightarrow \text{Al}^{2+}(\text{g}) + \text{e}^-$
 iii $\text{Ca}(\text{s}) + \frac{1}{2}\text{O}_2(\text{g}) \rightarrow \text{CaO}(\text{s})$
 iv $\text{Fe}(\text{s}) + 1\frac{1}{2}\text{Cl}_2(\text{g}) \rightarrow \text{FeCl}_3(\text{s})$
- b $\Delta H_{\text{form}}^{\ominus} = \Delta H_{\text{at}}^{\ominus} - [\Delta H_{\text{ion}}^{\ominus}(\text{Na}) + \Delta H_{\text{ion}}^{\ominus}(\text{Na}) + \Delta H_{\text{at}}^{\ominus}(\frac{1}{2}\text{Cl}_2(\text{g})) + \Delta H_{\text{at}}^{\ominus}(\text{Cl})]$
 $\Delta H_{\text{form}}^{\ominus} = (-411) - \{(+107) + (+496) + (+122) + (-348)\}$
 $= +377 \text{ kJ mol}^{-1}$
 $\Delta H_{\text{form}}^{\ominus} = (-411) - (+377) = -788 \text{ kJ mol}^{-1}$



- b i enthalpy change of atomisation of iodine
 ii first electron affinity of nitrogen
 iii enthalpy change of formation of strontium chloride
 iv lattice energy of cadmium chloride



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James F. Hall



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1. Understanding the eBook Chemistry Chapter 19 Assessment Answers
 - The Rise of Digital Reading Chemistry Chapter 19 Assessment Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chapter 19 Assessment Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Chapter 19 Assessment Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Chapter 19 Assessment Answers
 - Personalized Recommendations
 - Chemistry Chapter 19 Assessment Answers User Reviews and Ratings
 - Chemistry Chapter 19 Assessment Answers and Bestseller Lists

5. Accessing Chemistry Chapter 19 Assessment Answers Free and Paid eBooks
 - Chemistry Chapter 19 Assessment Answers Public Domain eBooks
 - Chemistry Chapter 19 Assessment Answers eBook Subscription Services
 - Chemistry Chapter 19 Assessment Answers Budget-Friendly Options
6. Navigating Chemistry Chapter 19 Assessment Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Chapter 19 Assessment Answers Compatibility with Devices
 - Chemistry Chapter 19 Assessment Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Chapter 19 Assessment Answers
 - Highlighting and Note-Taking Chemistry Chapter 19 Assessment Answers
 - Interactive Elements Chemistry Chapter 19 Assessment Answers
8. Staying Engaged with Chemistry Chapter 19 Assessment Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Chapter 19 Assessment Answers
9. Balancing eBooks and Physical Books Chemistry Chapter 19 Assessment Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Chapter 19 Assessment Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Chapter 19 Assessment Answers
 - Setting Reading Goals Chemistry Chapter 19 Assessment Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Chapter 19 Assessment Answers
 - Fact-Checking eBook Content of Chemistry Chapter 19 Assessment Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

-
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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