

has the K shown. If a reaction mixture has the concentrations as shown, in what direction is the reaction proceeding? Concentrations: $[\text{CO}_2] = 1.22 \times 10^{-2} \text{ M}$

$$[\text{CO}] = 2.00 \times 10^{-2} \text{ M} \quad [\text{C}_2\text{H}_2] = 3.00 \times 10^{-2} \text{ M}$$

- A) Right to left B) Left to Right
 C) It is at equilibrium D) Not enough information to determine

8. The K_{eq} for the equilibrium below is 0.112 at 700.0 °C.



What is the value of K_{eq} at this temperature for the following reaction?



- A) 0.224 B) 0.0425 C) 0.112 D) 6.93 E) -0.112

9. At 200 °C, the equilibrium constant (K_p) for the reaction $2\text{H}_2\text{O}(g) \rightleftharpoons \text{H}_2(g) + \text{O}_2(g)$ is 2.4×10^{-4} . A closed vessel is charged with 26.1 atm of H_2O . At equilibrium, the partial pressure of O_2 is _____ atm.

- A) 294 B) 35.7 C) 17.9 D) 6.00 E) 1.00×10^{-4}

10. In the coal-gasification process, carbon monoxide is converted to carbon dioxide via the following reaction: $\text{CO}(g) + \text{H}_2\text{O}(g) \rightleftharpoons \text{CO}_2(g) + \text{H}_2(g)$. In an experiment, 0.35 mol of CO and 0.40 mol of H_2O were placed in a 1.00-L reaction vessel. At equilibrium, there were 0.19 mol of CO remaining. K_{eq} at the temperature of the experiment is _____.

- A) 5.47 B) 0.75 C) 1.78 D) 0.56 E) 1.0

Short Answer:

1. Nitrosyl bromide decomposes according to the following equation.



A sample of NOBr (0.44 mol) was placed in a 2.00-L flask containing no NO or Br_2 . If the $K_{eq} = 56.2$, how many moles of NO and Br_2 , respectively, are in the flask at equilibrium? (4T, 1C)

2. The reaction: $2\text{H}_2\text{O}(g) + 3\text{O}_2(g) \rightleftharpoons 2\text{H}_2\text{O}_2(g)$ $K = 3.9 \times 10^4$

has the K shown. If a reaction mixture has the concentrations as shown,

Chemistry Chemical Equilibrium Test Answers

TD Snyder



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