

# Empirical and Molecular Formulas Worksheet

Key

## Objectives:

- be able to calculate empirical and molecular formulas

## Empirical Formula

1) What is the empirical formula of a compound that contains 0.783g of Carbon, 0.196g of Hydrogen and 0.521g of Oxygen?

$$\begin{array}{l}
 \text{C} \quad \frac{0.783\text{g C}}{12.01\text{g}} = \frac{0.065}{0.033} \approx 1.96 \quad 2 \\
 \text{H} \quad \frac{0.196\text{g H}}{1.01\text{g}} = \frac{0.194}{0.033} \approx 5.88 \quad 6 \\
 \text{O} \quad \frac{0.521\text{g O}}{16.00\text{g}} = \frac{0.033}{0.033} = 1 \quad 1
 \end{array}$$

$\text{C}_2\text{H}_6\text{O}$

2) What is empirical formula of a compound which consists of 89.14% Au and 10.80% of O?

$$\begin{array}{l}
 \text{Au} \quad \frac{89.14\text{g}}{196.97} = \frac{0.45}{0.45} = 1 \quad \text{AuO}_{1.5} \\
 \text{O} \quad \frac{10.80\text{g}}{16\text{g}} = \frac{0.675}{0.45} = 1.5
 \end{array}$$

$\text{Au}_2\text{O}_3$

3) What is empirical formula if compound consists of 21.2%N, 6.1%H, 24.2%S and 48.5%O?

$$\begin{array}{l}
 \text{N} \quad \frac{21.2\text{g}}{14.01} = 1.5 \quad \cdot \quad 2 \\
 \text{H} \quad \frac{6.1\text{g}}{1.01\text{g}} = 6.03 \quad \cdot \quad 8 \\
 \text{S} \quad \frac{24.2\text{g}}{32} = 0.75 \quad \cdot \quad 1 \\
 \text{O} \quad \frac{48.5\text{g}}{16.00\text{g}} = 3.03 \quad \cdot \quad 4
 \end{array}$$

$\text{N}_2\text{H}_8\text{S}_4\text{O}_4$

## Molecular Formula

4) Empirical formula of a substance is  $\text{CH}_2\text{O}$ . Molar mass is 180. What is the molecular formula?

$$\text{EFM} = (12 + 2 \times 1 + 16) = 30$$

$$\frac{180}{30} = 6 \quad \text{C}_6\text{H}_{12}\text{O}_6$$

5) Sample (3.585g) contains 1.388g of C, 0.345g of H, 1.850g O and its molar mass is 62g. What is molecular formula of this substance?

$$\begin{array}{l}
 \text{C} \quad \frac{1.388}{3.585} \rightarrow 38.7\% \rightarrow \frac{3.87\text{g}}{12.01\text{g}} = \frac{3.22}{3.22} = 1 \\
 \text{H} \quad \frac{0.345}{3.585} \rightarrow \frac{9.62\text{g}}{1.01\text{g}} = \frac{9.52}{3.22} = 3 \\
 \text{O} \quad \frac{1.850}{3.585} \rightarrow \frac{51.6\text{g}}{16\text{g}} = \frac{3.22}{3.22} = 1
 \end{array}$$

$\text{C}_2\text{H}_6\text{O}_2$

$\text{CH}_3\text{O} = \text{EFM} (12 + 3 \times 1 + 16)$

$$\frac{62}{31.03} = 2$$

1	$\text{C}_2\text{H}_6\text{O}$
2	$\text{Au}_2\text{O}_3$
3	$\text{NH}_4\text{SO}_4$
4	$\text{C}_2\text{H}_6\text{O}_2$
5	$\text{C}_2\text{H}_6\text{O}$

# Chemistry Empirical Formula Worksheet Answers

**M Tight**



## **Chemistry Empirical Formula Worksheet Answers:**

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**CONTENTS**

- 1 Writing the Formulae of Ionic Compounds
- 2 Formula Masses
- 3 Mass and Percent Compositions
- 4 The Mole and Mass
- 5 The Mole and Number of Particles
- 6 The Mole and Concentration
- 7 The Mole and Molar Volume
- 8 Empirical and Molecular Formulae
- 9 Chemical Equations
- 10 Calculations Based on Chemical Equations
- 11 Limiting Reagent
- 12 Gas Volume
- 13 Volume Stoichiometry
- 13 Percent Yield

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Jackson,2006-02-22 This 6 page study guide contains basic chemistry analysis and concepts designed specifically to aid science students

**Percent Composition and Empirical Formulas** ,2016 In this lecture you encounter two classic types of chemistry problems and learn the basic characteristics of each The lecture concludes with several practice problems to help you master the skill of solving percent composition problems

**Chemical Calculations** James Scott Long,Harold

Victor Anderson,1928

**Practice Makes Perfect Chemistry Review and Workbook, Second Edition** Marian DeWane,Heather Hattori,2018-12-28 The Winning Equation for Success in Chemistry is Practice Practice Practice This book will help you apply concepts and see how chemistry topics are interconnected Inside are numerous lessons to help you better understand the subject These lessons are accompanied by dozens of exercises to practice what you ve learned along with a complete answer key to check your work Throughout this book you will learn the terms to help you understand chemistry and you will expand your knowledge of the subject through hundreds of sample questions and their solutions With the lessons in this book you will find it easier than ever to grasp chemistry concepts And with a variety of exercises for practice you will gain confidence using your growing chemistry skills in your classwork and on exams

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## **Table of Contents Chemistry Empirical Formula Worksheet Answers**

1. Understanding the eBook Chemistry Empirical Formula Worksheet Answers
  - The Rise of Digital Reading Chemistry Empirical Formula Worksheet Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Empirical Formula Worksheet Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Empirical Formula Worksheet Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Empirical Formula Worksheet Answers
  - Personalized Recommendations
  - Chemistry Empirical Formula Worksheet Answers User Reviews and Ratings
  - Chemistry Empirical Formula Worksheet Answers and Bestseller Lists
5. Accessing Chemistry Empirical Formula Worksheet Answers Free and Paid eBooks
  - Chemistry Empirical Formula Worksheet Answers Public Domain eBooks
  - Chemistry Empirical Formula Worksheet Answers eBook Subscription Services
  - Chemistry Empirical Formula Worksheet Answers Budget-Friendly Options

6. Navigating Chemistry Empirical Formula Worksheet Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry Empirical Formula Worksheet Answers Compatibility with Devices
  - Chemistry Empirical Formula Worksheet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Empirical Formula Worksheet Answers
  - Highlighting and Note-Taking Chemistry Empirical Formula Worksheet Answers
  - Interactive Elements Chemistry Empirical Formula Worksheet Answers
8. Staying Engaged with Chemistry Empirical Formula Worksheet Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Empirical Formula Worksheet Answers
9. Balancing eBooks and Physical Books Chemistry Empirical Formula Worksheet Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Empirical Formula Worksheet Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Empirical Formula Worksheet Answers
  - Setting Reading Goals Chemistry Empirical Formula Worksheet Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Empirical Formula Worksheet Answers
  - Fact-Checking eBook Content of Chemistry Empirical Formula Worksheet Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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