

## Chemistry Unit 4 Test Review

### Electron Configuration

1. What are shapes of s, p, and d subshell?

s – sphere      p – dumbbell      d – clover leaf

2. Where are the s, p, d, and f subshell located on the periodic table?

s – group 1-2      p – group 13-18      d – group 3-12 (transition metals)

3. How many electrons are needed to completely fill the following energy levels?

a. 1<sup>st</sup> energy level – 2 electrons

b. 2<sup>nd</sup> energy level – 8 electrons

c. 3<sup>rd</sup> energy level – 18 electrons

4. Which one of the three rules explains why electrons have opposite spins?

Pauli exclusion principle

5. Which rule tells us that one electron must fill each orbital before pairing up?

Hund's rule

6. Which rule explains why we must fill lowest energy levels first?

Aufbau principle

7. Complete the orbital notation, electron configuration, and noble gas configuration for the following element at ground state.

a. Ca



b. P



c. Cr



# Chemistry Unit 4 Test Review Electron Configuration

**Patrick Vollmar**



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