



Chapter Test Answers

CHAPTER 1: SOLUTIONS

Ans. 1. (a) 2. (d) 3. (a) 4. (c) 5. (a) 6. (b)

Ans. 7. (i) Acetone and aniline mixture represents negative deviations from Raoult's law, hence for this mixture, ΔH_{mix} and ΔV_{mix} is negative.

(ii) For positive deviations,

$$p_A > p_A^0 x_A \text{ and } p_B > p_B^0 x_B$$

(iii)



The above graph is correct for water and nitric acid mixture.

Water and nitric acid mixture shows negative deviations from Raoult's law, hence

$$p_A < p_A^0 x_A \text{ and } p_B < p_B^0 x_B$$

Or

Water-HCl mixture shows negative deviations from Raoult's law and solutions showing negative deviations from ideal behaviour form maximum boiling azeotrope.

Ans. 8. Gas B will have the higher value of K_{H_2} (Henry's Constant) as lower is the solubility of the gas in the liquid higher is the value of K_{H_2} .

Ans. 9. In case of liquids, due to increase of pressure inside the cooker, the boiling point of water increases leading to faster cooking than in pan.

Ans. 10.

Ideal solution	Non-ideal solution
The components of this solution obey Raoult's law at all temperatures and concentrations. i.e. $p_A = p_A^0 x_A$ and $p_B = p_B^0 x_B$	The components of this solution do not obey Raoult's law. They show positive and negative deviations from Raoult's law. i.e. $p_A \neq p_A^0 x_A$ and $p_B \neq p_B^0 x_B$
On mixing, there is no enthalpy and volume change. i.e. $\Delta H_{mix} = 0$ and $\Delta V_{mix} = 0$	On mixing, there is enthalpy and volume change. i.e. $\Delta H_{mix} \neq 0$ and $\Delta V_{mix} \neq 0$

Ans. 11.



TIP

(i) First find out ΔT_f by using the formula

$$\Delta T_f = \frac{K_f \times W_2 \times 1000}{M_2 \times W_1}$$

(ii) Then, find out the value of T_f by using formula,

$$\Delta T_f = T_f^0 - T_f$$

Given, W_2 (glucose) = 0.520 g

W_1 (H_2O) = 80.2 g

K_f (H_2O) = 1.86 Km^{-1}

M_2 of $C_6H_{12}O_6$ (glucose) = 180 g mol^{-1}

$$\Delta T_f = \frac{K_f \times W_2 \times 1000}{M_2 \times W_1}$$

$$= \frac{1.86 \text{ Km}^{-1} \times 0.520 \text{ g} \times 1000}{180 \text{ g mol}^{-1} \times 80.2 \text{ g}}$$

$$= 0.0669 \text{ K}$$

Freezing point of solution

$$T_f = 273 \text{ K} - 0.0669 \text{ K}$$

$$= 272.933 \text{ K}$$

Ans. 12. van't Hoff factor is the ratio of the experimental value of colligative property to the calculated value of the colligative property. It is used to find out the extent of dissociation or association.

In dissociation, observed molar mass has lesser value than normal, so i is more than one ($i > 1$), e.g. The value of i for ethanoic acid in benzene is nearly 0.5, while for aqueous KCl, NaCl and $MgSO_4$ solution, it is nearly 2. The value of i for K_2SO_4 is nearly 3.

Ans. 13. (i) $\Delta T_f = \frac{K_f W_2 \times 1000}{M_2 \times W_1}$

$$0.383 = \left(\frac{3.83 \times 2.56}{M \times 100} \right) \times 100$$

$$M = 256$$

$$5 \times x = 256$$

$$132 \times x = 256$$

$$x = 8$$

(ii) (a) If we place cells in solution containing 12% sodium chloride solution, they shrink.

(b) If we place blood cells in solution containing 0.4% sodium chloride solution, they swell.

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