

To receive full credit, show all work with units for each question. Refer to the table below for appropriate constants, boiling points, etc.

Compound	Formula	Density (20 °C, g cm ⁻³)	FP (°C)	BP (°C)	K _f (°C/m)	K _b (°C/m)
Acetic Acid	CH ₃ COOH	1.0492	16.6	117.9	-3.90	3.07
Benzene	C ₆ H ₆	0.8765	5.5	80.1	-4.90	2.53
Carbon Tetrachloride	CCl ₄	1.5940	-22.99	76.54	-2.98	5.03
Nitrobenzene	C ₆ H ₅ NO ₂	1.2037	5.7	210.8	-7.00	5.24
Water	H ₂ O	0.998203	0.000	100.0	-1.86	0.512
Ethanol	C ₂ H ₅ O	0.7893	-117.3	78.5	-1.99	1.22

- For a mixture prepared by mixing 5.00 g of caffeine (C₈H₁₀N₄O₂) with 250.0 g of water
 - Calculate the boiling point of this mixture
 - Calculate the freezing point of this mixture
- For a mixture prepared by mixing 10.0 g of sodium sulfate in 250 g of water
 - Calculate the boiling point of this mixture
 - Calculate the freezing point of this mixture
- You are examining forensics data for a court case and need to identify an analgesic (pain killer). It could be either acetylsalicylic acid (HC₉H₇O₄, 180.160 g/mol), ibuprofen (C₁₃H₁₈O₂, 206.3 g/mol) or acetaminophen (C₈H₉NO₂, 151.17 g/mol). A mixture of the unknown is prepared by mixing 10.0 grams of unknown and diluting to 250 mL with ethanol. The boiling point of this mixture is 78.8 °C. The density of this solution is 0.80 g/mL. Assume that the unknown takes up no volume in the solution. Determine the identity of the unknown. Big hint: Can you calculate the molar mass?

Calculations Involving Colligative Properties Section Review Answers

**Gary Leighton Long, Sara Long, Robert
C. Fay**

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