

Calculus II
Practice Problems 1: Answers

1. Solve for x :

a) $6^x = 36^{2-x}$

Answer. Since $36 = 6^2$, the equation becomes $6^x = 6^{2(2-x)}$, so we must have $x = 2(2-x)$ which has the solution $x = 4/3$.

b) $\ln_3 x = 5$

Answer. If we exponentiate both sides we get $x = 3^5 = 243$.

c) $\ln_2(x+1) - \ln_2(x-1) = \ln_2 8$

Answer. Since the difference of logarithms is the logarithm of the quotient, we rewrite this as

$$\ln_2\left(\frac{x+1}{x-1}\right) = \ln_2 8,$$

which is, after exponentiating, the same as

$$\frac{x+1}{x-1} = 8,$$

which gives us $x+1 = 8(x-1)$, so that $x = 9/7$.

2. Find the derivative of the given function:

a) $y = \ln(\ln x)$

Answer. Use the chain rule:

$$\frac{dy}{dx} = \frac{1}{\ln x} \frac{d}{dx} \ln x = \frac{1}{x \ln x}$$

b) $y = \log_2(x^2 + 1)$

Answer. Remember that $\log_2 A = \ln A / \ln 2$, so $y = (\ln(x^2 + 1)) / \ln 2$. Then, use the chain rule:

$$\frac{dy}{dx} = \frac{1}{(\ln 2)(x^2 + 1)} 2x = \frac{2}{\ln 2} \frac{x}{x^2 + 1}.$$

c) $y = \frac{e^{x^2}}{x}$

Answer. Use the quotient rule carefully:

$$\begin{aligned} \frac{dy}{dx} &= \frac{x(2xe^{x^2}) - e^{x^2}}{x^2} = 2e^{x^2} - x^{-2}e^{x^2} \\ &= e^{x^2}(2 - x^{-2}) \end{aligned}$$

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Lauren Gardner



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for the nitration 4 01 ml of nitric acid 4 00 ml of sulfuric acid and 4 56 ml bromobenzene were used to complete the reaction with these measurements bromobenzene was found to be the limiting reagent since only 4 56 ml of bromobenzene was used the amount of product predicted to be recovered was 8 7533 g

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in this experiment we will carry out the nitration of bromobenzene bromine substituted on an aromatic ring is an ortho para directing group it will direct an incoming electrophile to the ortho and para positions it is also deactivating meaning that its reactions will be slower than the identical reactions of benzene

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section 20711 lab experiment nitration of bromobenzene introduction the chemical reaction of nitration consists of a nitro group being added to or substituted in a molecule nitration can basically be carried out by a mixture of concentrated nitric acid and sulfuric acid this mixture is useful to obtain the active nitronium ion

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1 go to the restroom and wash the affected area with lots of soap and water the soap in the lab is not suitable for this purpose 2 return to the lab and apply mineral oil to the affected area 3 the summary to this warning is not to touch any part of your body while performing this experiment if you must leave the lab for any reason 1

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in the nitration of bromobenzene a mixture of sulfuric acid and nitric acid react with one another nitric acid the weaker of the two acids will serve as base to form nitronium ion bromobenzene will attack the nitronium ion by nucleophilic attack

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then actual is your actual findings yield actual theoretical $\times 100$ what are the possible side reactions of nitration of bromobenzene dinitration can occur at ortho para or at both ortho positions occurs in overheating meta product happens very rarely because bromine is an ortho para director

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the soap in the lab is not suitable for this purpose 2 return to the lab and apply mineral oil to the affected area theory and explanation of the nitration of bromobenzene if the temperature of the reaction mixture exceeds 60 o

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the bromine substituent in bromobenzene is an ortho para director the para position is favored due to steric effects in the reaction that takes place in this lab a nitro group is substituted onto bromobenzene at the para position to form 1 bromo 4 nitrobenzene

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over 5 minutes slowly pipet this bromobenzene into the icy cold nitration mixture hno₃ h₂so₄ in the 125 ml flask and stir continue to stir the reaction mixture in the ice bath for 30 minutes do not heat in a new 250 ml erlenmeyer

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on the right the arenium cation has a nitro set with incredibly small 0.6 angstrom bond distances this will result in an optimization failure as the atoms are even close for the programme to accomodate their electrons demonstration of creating the p arenium cation by to nitration of bromobenzene

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the purpose of this lab was to observe where substituents are added on a benzene ring by way of electrophilic aromatic substitution the electrophilic source in the nitration of bromobenzene comes from a nitronium ion no₂⁺ which acts as the electrophilic species

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