

### Basic Stoichiometry Post-Lab Homework Exercises

1. Load the "Reactants, Products, and Leftovers" simulation and work through each of the levels of the Game! At <http://phet.colorado.edu/> or googling "phet." You may have to download or update Java on your computer.

**Show your work for all calculations. Remember: Labels, units, sig figs and cancellation!**

2. For the reaction  $P_4 + 6Cl_2 \rightarrow 4PCl_3$ , determine how many moles of chlorine  $Cl_2$  would be needed to react with 4.0 moles of phosphorus  $P_4$  to entirely use up all the phosphorus. 4) \_\_\_\_\_

3. If 3.00 moles of  $P_4$  reacted with 22.0 moles  $Cl_2$  according to the above reaction, determine:

- a) How many moles  $PCl_3$  are produced a) \_\_\_\_\_
- b) How many moles of  $P_4$  are left in excess after the reaction (if any) b) \_\_\_\_\_
- c) How many moles of  $Cl_2$  are left in excess after the reaction (if any) c) \_\_\_\_\_

In reality, reactants don't have to react in perfect whole-numbers of moles. In a two-reactant synthesis reaction, usually **one reactant gets entirely used up (and determines how much product is made)**, even if that means using fractions of a mole of reactant. For instance, when solid, metallic aluminum  $Al$  and red, liquid bromine  $Br_2$  are brought together, they make a white solid according to the reaction  $2Al + 3Br_2 \rightarrow 2AlBr_3$ . If 5.0 moles of aluminum  $Al$  was reacted with 10 moles bromine  $Br_2$ , all five moles of aluminum would react, with only 7.5 moles bromine. (2:3 mole ratio) This would produce only 5.0 moles of  $AlBr_3$ , leaving 2.5 moles of excess  $Br_2$  behind.

5. Now assume 3.0 moles  $Al$  and 4.0 moles  $Br_2$  react.

- a) How much (in moles)  $AlBr_3$  gets produced? a) \_\_\_\_\_
- b) Which chemical must be the *excess reactant*? b) \_\_\_\_\_
- c) Which chemical is the *limiting reactant*? c) \_\_\_\_\_
- d) If all the limiting reactant gets used up, how much of the excess reactant is left? d) \_\_\_\_\_

# Basic Stoichiometry Post Lab Homework Exercises

## Answer Key

**S Ashworth**



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## **Table of Contents Basic Stoichiometry Post Lab Homework Exercises Answer Key**

1. Understanding the eBook Basic Stoichiometry Post Lab Homework Exercises Answer Key
  - The Rise of Digital Reading Basic Stoichiometry Post Lab Homework Exercises Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Basic Stoichiometry Post Lab Homework Exercises Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Basic Stoichiometry Post Lab Homework Exercises Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Basic Stoichiometry Post Lab Homework Exercises Answer Key
  - Personalized Recommendations
  - Basic Stoichiometry Post Lab Homework Exercises Answer Key User Reviews and Ratings
  - Basic Stoichiometry Post Lab Homework Exercises Answer Key and Bestseller Lists
5. Accessing Basic Stoichiometry Post Lab Homework Exercises Answer Key Free and Paid eBooks

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6. Navigating Basic Stoichiometry Post Lab Homework Exercises Answer Key eBook Formats
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  - Basic Stoichiometry Post Lab Homework Exercises Answer Key Compatibility with Devices
  - Basic Stoichiometry Post Lab Homework Exercises Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Basic Stoichiometry Post Lab Homework Exercises Answer Key
  - Highlighting and Note-Taking Basic Stoichiometry Post Lab Homework Exercises Answer Key
  - Interactive Elements Basic Stoichiometry Post Lab Homework Exercises Answer Key
8. Staying Engaged with Basic Stoichiometry Post Lab Homework Exercises Answer Key
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Basic Stoichiometry Post Lab Homework Exercises Answer Key
9. Balancing eBooks and Physical Books Basic Stoichiometry Post Lab Homework Exercises Answer Key
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Basic Stoichiometry Post Lab Homework Exercises Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Basic Stoichiometry Post Lab Homework Exercises Answer Key
- Setting Reading Goals Basic Stoichiometry Post Lab Homework Exercises Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Basic Stoichiometry Post Lab Homework Exercises Answer Key
- Fact-Checking eBook Content of Basic Stoichiometry Post Lab Homework Exercises Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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