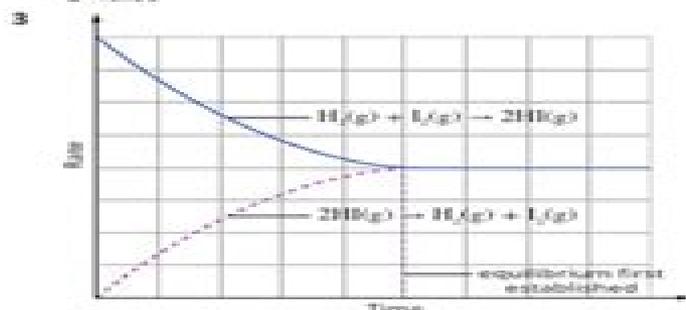


## Chapter 2 Chemical equilibrium

### 2.1 Dynamic equilibrium

#### 2.1 KEY QUESTIONS

- An **open** system occurs when both matter and energy are exchanged between the system and the surroundings.
  - Chemical equilibria can occur in **closed** systems.
  - In some circumstances, a reversible reaction can form a **dynamic** equilibrium.
  - The extent of reaction for chemical equilibria is **different** for different equilibria.
- false
  - false
  - true
  - false



- closed, decreases, less, reverse, increases, equal
- open; both energy and matter (carbon dioxide and water vapour) are exchanged with the surroundings
  - closed; only energy can be exchanged with the surroundings. While the can is unopened, no matter is exchanged.
  - open; both energy and matter are exchanged with the surroundings
- B and F. Equilibrium can only occur in a closed system. At the molecular level, some of the sugar in the saturated sugar solution will continue to dissolve at the same rate as some is precipitated. In the unopened bottle of sparkling mineral water, the dissolved carbon dioxide will be in equilibrium with the gaseous carbon dioxide. In all of the other situations, the reverse reaction is not occurring at the same rate as the forward reaction, because the systems are open.
- $130\text{ kJ mol}^{-1}$ . The activation energy is the difference between  $390\text{ kJ mol}^{-1}$  and  $260\text{ kJ mol}^{-1}$ , which is  $130\text{ kJ mol}^{-1}$ .
- Chemical equilibrium is 'dynamic' because both forward and reverse reactions occur at the same rate. An equilibrium develops between water vapour and water when wet clothes are in a sealed bag, with water evaporating as rapidly as water vapour condenses, so the clothes remain wet.
  - When the bag is opened, water vapour escapes and the rate of evaporation exceeds the rate of condensation. The system is not in equilibrium and the clothes dry.
- $0.07\text{ mol L}^{-1}$
  - $0\text{ mol L}^{-1}$
  - $0.03\text{ mol L}^{-1}$
  - $0.08\text{ mol L}^{-1}$
  - $0.04\text{ moles}$
  - The horizontal regions of the graph indicate that there is no change in concentration of  $\text{NO}_2$  or  $\text{N}_2\text{O}_4$ . This is when the system is at equilibrium and both the forward and reverse reactions are occurring at the same rate.
  - 6 seconds
  - As the reaction proceeds, the intensity of the brown colour will increase as the concentration of  $\text{NO}_2$  increases until equilibrium is reached. Once equilibrium is reached, the colour of the reaction mixture will remain constant.

# Pearson Reaction Rates And Equilibrium Packet Answers

**E Durkheim**



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