

Chem 225 Analytical Chemistry 1

Problem Set No. 1_2

Redox Titrations

1. Selenium in a 10.0-g soil sample is distilled as the tetrabromide, which is collected in aqueous solution where it is hydrolyzed to SeO_3^{2-} . The SeO_3^{2-} is determined iodometrically, requiring 4.5 mL of standard thiosulfate solution for the titration. If the thiosulfate titer is 0.049 mg $\text{K}_2\text{Cr}_2\text{O}_7/\text{mL}$, what is the concentration of selenium in the soil in ppm?
2. The calcium in a 5.00-mL serum sample is precipitated as CaC_2O_4 with ammonium oxalate. The filtered precipitate is dissolved in acid, the solution is heated, and the oxalate is titrated with 0.00100 M KMnO_4 requiring 4.94 mL. Calculate the concentration of calcium in the serum in meq/L (equivalents based on charge).
3. A 2.50-g sample containing As_2O_3 , Na_2HASO_3 , and inert material is dissolved, and the pH is adjusted to neutral with excess NaHCO_3 . The As(III) is titrated with 0.150 M I_2 solution, requiring 11.3 mL to just reach the end point. Then, the solution (all the arsenic in the +5 state now) is acidified with HCl, excess KI is added, and the liberated I_2 is titrated with 0.120 M $\text{Na}_2\text{S}_2\text{O}_3$, requiring 41.2 mL. Calculate the percent As_2O_3 and Na_2HASO_3 in the sample.
4. If 1.00 mL KMnO_4 solution will react with 0.125 g Fe^{2+} and if 1.00 mL $\text{KHC}_2\text{O}_4 \cdot \text{H}_2\text{C}_2\text{O}_4$ solution will react with 0.175 mL of the KMnO_4 solution, how many milliliters of 0.200 M NaOH will react with 1.00 mL of the tetroxalate solution? (All three protons on the tetroxalate are titratable.)
5. The sulfide content in a pulp plant effluent is determined with a sulfide ion- selective electrode using the method of standard additions for calibration. A 10.0-mL sample is diluted to 25.0 mL with water and gives a potential reading of -216.4 mV. A similar 10.0-mL sample plus 1.00 mL of 0.030 M sulfide standard diluted to 25.0 mL gives a reading of -224.0 mV. Calculate the concentration of sulfide in the sample.

Redox Problem Set 1 Reactions And Stoichiometry

RJ Alexander



Redox Problem Set 1 Reactions And Stoichiometry:

Redox Problem Set 1 Reactions And Stoichiometry Book Review: Unveiling the Magic of Language

In an electronic era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**Redox Problem Set 1 Reactions And Stoichiometry**," compiled by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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