

CHEMISTRY



Properties of Matter

- Matter is anything that has mass and occupies space.
- Matter can be classified into solids, liquids, and gases.
- Solids have a definite shape and volume.
- Liquids have a definite volume but no definite shape.
- Gases have neither a definite shape nor a definite volume.

States of Matter

- Solids: Particles are closely packed and vibrate about their mean positions.
- Liquids: Particles are less closely packed and can move past one another.
- Gases: Particles are far apart and move randomly in all directions.

Atomic Structure

Atoms are the smallest particles that cannot be created or destroyed in a chemical reaction. They are made up of three sub-particles: electrons, protons, and neutrons.

Electrons

- Discovered by Cathode ray experiment.
- Charge: -1.6×10^{-19} C
- Mass: 9.1×10^{-31} kg

Protons

- Discovered by Rutherford's alpha-particle scattering experiment.
- Charge: $+1.6 \times 10^{-19}$ C
- Mass: 1.67×10^{-27} kg

Neutrons

- Discovered by Chadwick.
- Charge: 0
- Mass: 1.67×10^{-27} kg

Atomic Structure

Atoms are made up of three sub-particles: electrons, protons, and neutrons. The nucleus is made up of protons and neutrons. Electrons are present in shells around the nucleus.

Atomic Number

The atomic number of an element is the number of protons present in the nucleus of an atom of that element.

Mass Number

The mass number of an element is the sum of the number of protons and neutrons present in the nucleus of an atom of that element.

Element	Atomic Number (Z)	Mass Number (A)
Hydrogen	1	1
Helium	2	4
Lithium	3	7
Boron	5	11
Carbon	6	12
Nitrogen	7	14
Oxygen	8	16
Fluorine	9	19
Neon	10	20

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