

Acid-Base Titration Lab Report Sheet

Name:	Quiz is 20 points
	Lab Report is 80 points

Part I: Standardization of NaOH using Anhydrous Citric Acid

	Trial 1	Trial 2	Trial 3
Starting M of citric acid solution	$8.328 \times 10^{-3} \text{ M}$	$8.328 \times 10^{-3} \text{ M}$	$8.328 \times 10^{-3} \text{ M}$
Volume of citric acid	20.00 mL	20.00 mL	20.00 mL
Moles of citric acid	$1.667 \times 10^{-4} \text{ mol}$	$1.667 \times 10^{-4} \text{ mol}$	$1.667 \times 10^{-4} \text{ mol}$
Moles of NaOH	$4.997 \times 10^{-4} \text{ mol}$	$4.997 \times 10^{-4} \text{ mol}$	$4.997 \times 10^{-4} \text{ mol}$
Initial graduated cylinder volume	10.0 mL	10.0 mL	10.0 mL
Final graduated cylinder volume	5.2 mL	5.4 mL	5.3 mL
Volume of NaOH used	4.8 mL	4.6 mL	4.7 mL
Volume of NaOH used in L	0.0048 L	0.0046 L	0.0047 L
NaOH concentration	0.104 M	0.109 M	0.106 M

Average Molarity of NaOH _____ 0.106 M _____

Show your work for one trial as well as the average molarity.

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SJ Ball



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