



COVALENT BONDING

Practice Problems

Do your answers to these very practice problems.

SECTION 8.1 MOLECULAR COMPOUNDS

1. Classify each of the following as ionic or molecular.

- a. NaCl (aq) d. H_2SO_4 (aq) g. HCl (aq)
 b. H_2O (l) e. NaCl (solid) h. H_2O (solid)

2. Which of the following is a molecular compound?

- a. H_2O d. H_2O_2
 b. NaCl e. Na_2O f. H_2O_2

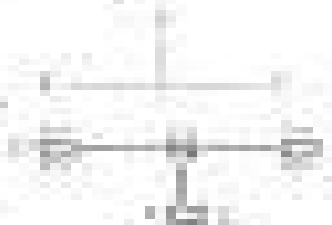
3. Which of the following is a molecular compound? Write the formula for each.
 a. H_2O (l) b. H_2O_2 (l) c. NaCl (aq) d. H_2O (solid)
 e. H_2O_2 (solid) f. NaCl (solid) g. H_2O (solid) h. H_2O_2 (solid)

SECTION 8.2 THE NATURE OF COVALENT BONDING

1. Draw the Lewis structure for H-F .



2. Draw the Lewis structure for CH_4 .



3. Draw the Lewis structure for CH_3Cl .



4. Draw the Lewis structure for $\text{H-C}\equiv\text{C-H}$.

5. Draw the Lewis structure for H_2O . Show the lone pairs on oxygen.



6. Draw the Lewis structure for H_2O_2 .

SECTION 8.3 BONDING THE ORBITALS

1. Which orbital is the most negative in H_2O ? Which orbital is the most positive in H_2O ?
 2. Which orbital is the most negative in H_2O ? Which orbital is the most positive in H_2O ?

Chapter 8 Covalent Bonding Practice Problems Answers

JA Banks



Chapter 8 Covalent Bonding Practice Problems Answers:

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Table of Contents Chapter 8 Covalent Bonding Practice Problems Answers

1. Understanding the eBook Chapter 8 Covalent Bonding Practice Problems Answers
 - The Rise of Digital Reading Chapter 8 Covalent Bonding Practice Problems Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 8 Covalent Bonding Practice Problems Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 8 Covalent Bonding Practice Problems Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 8 Covalent Bonding Practice Problems Answers
 - Personalized Recommendations
 - Chapter 8 Covalent Bonding Practice Problems Answers User Reviews and Ratings
 - Chapter 8 Covalent Bonding Practice Problems Answers and Bestseller Lists
5. Accessing Chapter 8 Covalent Bonding Practice Problems Answers Free and Paid eBooks
 - Chapter 8 Covalent Bonding Practice Problems Answers Public Domain eBooks
 - Chapter 8 Covalent Bonding Practice Problems Answers eBook Subscription Services
 - Chapter 8 Covalent Bonding Practice Problems Answers Budget-Friendly Options

6. Navigating Chapter 8 Covalent Bonding Practice Problems Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 8 Covalent Bonding Practice Problems Answers Compatibility with Devices
 - Chapter 8 Covalent Bonding Practice Problems Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 8 Covalent Bonding Practice Problems Answers
 - Highlighting and Note-Taking Chapter 8 Covalent Bonding Practice Problems Answers
 - Interactive Elements Chapter 8 Covalent Bonding Practice Problems Answers
8. Staying Engaged with Chapter 8 Covalent Bonding Practice Problems Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 8 Covalent Bonding Practice Problems Answers
9. Balancing eBooks and Physical Books Chapter 8 Covalent Bonding Practice Problems Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 8 Covalent Bonding Practice Problems Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter 8 Covalent Bonding Practice Problems Answers
 - Setting Reading Goals Chapter 8 Covalent Bonding Practice Problems Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 8 Covalent Bonding Practice Problems Answers
 - Fact-Checking eBook Content of Chapter 8 Covalent Bonding Practice Problems Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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