

**CHEMFILE MINI-GUIDE TO PROBLEM SOLVING****2. PLAN**

- *What step is needed to calculate the concentration of the diluted solution?*

Apply the principle that  $\text{volume}_1 \times \text{molarity}_1 = \text{volume}_2 \times \text{molarity}_2$ .

$$M_1 V_1 = M_2 V_2 \xrightarrow[\text{substitute each of the known quantities for its symbol and calculate}]{\begin{array}{l} \text{rearrange the equation} \\ M_1 V_1 = M_2 V_2 \\ \text{algebraically to solve for } M_2 \end{array}} M_2 = \frac{M_1 V_1}{V_2}$$

$$M_2 = \frac{4.74 \text{ M} \times 50.00 \text{ mL}}{250.00 \text{ mL}}$$

**3. COMPUTE**

Note: Even though molarity is moles per liter, you can use volumes in milliliters along with molarity whenever the units cancel.

$$M_2 = \frac{4.74 \text{ M} \times 50.00 \text{ mL}}{250.00 \text{ mL}} = 0.948 \text{ M}$$

**4. EVALUATE**

- *Are the units correct?*
- *Is the number of significant figures correct?*
- *Is the answer reasonable?*

Yes; molarity (mol/L) was required.  
 Yes; the number of significant figures is correct because the data had a minimum of three significant figures.  
 Yes; the computation is the same as  $4.74/5$ , which is a little less than 1.

**PRACTICE**

1. Complete the table below by calculating the missing value in each row.

|    | <b>Molarity of stock solution</b> | <b>Volume of stock solution</b> | <b>Molarity of dilute solution</b> | <b>Volume of dilute solution</b> |              |
|----|-----------------------------------|---------------------------------|------------------------------------|----------------------------------|--------------|
| a. | 0.500 M KBr                       | 20.00 mL                        | 7 M KBr                            | 100.00 mL                        | ans: 0.100 M |
| b. | 1.00 M LiOH                       | ? mL                            | 0.075 M LiOH                       | 500.00 mL                        | ans: 38 mL   |
| c. | 7 M HI                            | 5.00 mL                         | 0.0493 M HI                        | 100.00 mL                        | ans: 0.986 M |
| d. | 12.0 M HCl                        | 0.250 L                         | 1.8 M HCl                          | ? L                              | ans: 1.7 L   |
| e. | 7.44 M NH <sub>3</sub>            | ? mL                            | 0.093 M NH <sub>3</sub>            | 4.00 L                           | ans: 50. mL  |

# Chemfile Mini Guide To Problem Solving Answer Key

**Raffaela Di Napoli**



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