

# Chemistry

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(Chapter 1)(The Solid State)  
(Class 12)  
Exercises

## Question 1.1:

Define the term 'amorphous'. Give a few examples of amorphous solids.

### Answer 1.1:

Amorphous solids are the solids whose constituent particles are of irregular shapes and have short range order. These solids are isotropic in nature and melt over a range of temperature. Therefore, amorphous solids are sometimes called pseudo solids or super cooled liquids. They do not have definite heat of fusion. When cut with a sharp-edged tool, they cut into two pieces with irregular surfaces. Examples of amorphous solids include glass, rubber, and plastic.

## Question 1.2:

What makes a glass different from a solid such as quartz? Under what conditions could quartz be converted into glass?

### Answer 1.2:

The arrangement of the constituent particles makes glass different from quartz. In glass, the constituent particles have short range order, but in quartz, the constituent particles have both long range and short range orders. Quartz can be converted into glass by heating and then cooling it rapidly.

## Question 1.3:

Classify each of the following solids as ionic, metallic, molecular, network (covalent) or amorphous.

- |  |                |
|--|----------------|
| (i) Tetra phosphorus deoxide ( $P_4O_{10}$ ) | (vii) Graphite |
| (ii) Ammonium phosphate $(NH_4)_3PO_4$       | (viii) Brass   |
| (iii) SiC                                    | (ix) Rb        |
| (iv) I <sub>2</sub>                          | (x) LiBr       |
| (v) P <sub>4</sub>                           | (xi) Si        |
| (vi) Plastic                                 |                |

### Answer 1.3:

- |                    |   |
|--------------------|---|
| Ionic              | → (ii) Ammonium phosphate $(NH_4)_3PO_4$ , (x) LiBr                                     |
| Metallic           | → (viii) Brass, (ix) Rb   |
| Molecular          | → (i) Tetra phosphorus deoxide $(P_4O_{10})$ , (iv) I <sub>2</sub> , (v) P <sub>4</sub> |
| Covalent (network) | → (iii) SiC, (vii) Graphite, (xi) Si  |
| Amorphous          | → (vi) Plastic  |

## Question 1.4:

- (i) What is meant by the term 'coordination number'?
- (ii) What is the coordination number of atoms:
- (a) in a cubic close-packed structure?
- (b) in a body-centred cubic structure?

### Answer 1.4

- (i) The number of nearest neighbours of any constituent particle present in the crystal lattice is called its coordination number.
- (ii) The coordination number of atoms
- (a) in a cubic close-packed structure is 12, and
- (b) in a body-centred cubic structure is 8

## Question 1.5:

How can you determine the atomic mass of an unknown metal if you know its density and the dimension of its unit cell? Explain.

### Answer 1.5:

By knowing the density of an unknown metal and the dimension of its unit cell, the atomic mass of the metal can be determined.

Let 'a' be the edge length of a unit cell of a crystal, 'd' be the density of the metal, 'm' be the atomic mass of the metal and 'Z' be the number of atoms in the unit cell.

Now, density of the unit cell =  $\frac{\text{Mass of the unit cell}}{\text{Volume of the unit cell}}$

$$\Rightarrow d = \frac{Zm}{a^3} \quad \dots (i)$$

[Since mass of the unit cell = Number of atoms in the unit cell × Atomic mass]

[Volume of the unit cell = (Edge length of the cubic unit cell)<sup>3</sup>]

# Chemistry Chapter 1

**Pharmacy Department of Thakur  
Publication**



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Principles Of Descriptive Inorganic Chemistry Gary Wulfsberg, 1991-05-29 This is the only text currently available organized by class of compound and by property or reaction type not group by group or element by element which requires students to memorize isolated facts This is the only text currently available organized by class of compound and by property or reaction type not group by group or element by element which requires students to memorize isolated facts Translated into Italian *Teaching Guide to Science and Cancer ... for the National Science Teachers Association* National Institutes of Health (U.S.), 1966 **D. Pharm Exit Examination (DPEE) Kit** Pharmacy Department of Thakur Publication, 2024-09-01 The D Pharm Exit Exam Kit by Thakur Publication is an essential study resource for students preparing for their D Pharm exit exams This comprehensive kit includes a wide range of practice questions solved papers and exam oriented content designed to help students revise and test their knowledge effectively AS PER PCI SYLLABUS 7000 MCQs COVERED ALL SUBJECTS With its user friendly format and reliable content the D Pharm Exit Exam Kit ensures students are well equipped to excel in their exams and embark on a successful pharmaceutical career Public Health Service Publication, 1966 **Science and Cancer** Ralph P. Frazier, 1966 **The Pearson Complete Guide For Aieee 2/e** Khattar, *Organic Synthesis* Michael Smith, 2016-11-22 *Organic Synthesis Fourth Edition* provides a reaction based approach to this important branch of organic chemistry Updated and accessible this eagerly awaited revision offers a comprehensive foundation for graduate students coming from disparate backgrounds and knowledge levels to provide them with critical working knowledge of basic reactions stereochemistry and conformational principles This reliable resource uniquely incorporates molecular modeling content problems and visualizations and includes reaction examples and homework problems drawn from the latest in the current literature In the Fourth Edition the organization of the book has been improved to better serve students and professors and accommodate important updates in the field The first chapter reviews basic retrosynthesis conformations and stereochemistry The next three chapters provide an introduction to and a review of functional group exchange reactions these are followed by chapters reviewing protecting groups oxidation and reduction reactions and reagents hydroboration selectivity in reactions A separate chapter discusses strategies of organic synthesis and the book then delves deeper in teaching the reactions required to actually complete a synthesis Carbon carbon bond formation reactions using both nucleophilic carbon reactions are presented and then electrophilic carbon reactions followed by pericyclic reactions and radical and carbene reactions The important organometallic reactions have been consolidated into a single chapter Finally the chapter on combinatorial chemistry has been removed from the strategies chapter and placed in a separate chapter along with valuable and forward looking content on green organic chemistry process chemistry and continuous flow chemistry Throughout the text *Organic Synthesis Fourth Edition* utilizes Spartan generated molecular models class tested content and useful pedagogical features to aid student study and retention including Chapter Review Questions and Homework Problems

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**Chemistry** Stanley R. Radel,Marjorie H. Navidi,1990      **Industrial & Engineering Chemistry** ,1924      *Schaums Outline of Organic Chemistry 5/E (ENHANCED EBOOK)* Herbert Meislich,Howard Nechamkin,Jacob Sharefkin,George J. Hademenos,2013-05-31 Tough Test Questions Missed Lectures Not Enough Time Fortunately there s Schaum s This all in one package includes more than 1 800 fully solved problems examples and practice exercises to sharpen your problem solving skills Plus you will have access to 24 detailed videos featuring Chemistry instructors who explain the most commonly tested concepts it s just like having your own virtual tutor You ll find everything you need to build confidence skills and knowledge for the highest score possible More than 40 million students have trusted Schaum s to help them succeed in the classroom and on exams Schaum s is the key to faster learning and higher grades in every subject Each Outline presents all the essential course information in an easy to follow topic by topic format You also get hundreds of examples solved problems and practice exercises to test your skills This Schaum s Outline gives you 1 806 fully solved problems Hundreds of examples with explanations of organic chemistry concepts Support for all the major textbooks for organic chemistry courses Fully compatible with your classroom text Schaum s highlights all the important facts you need to know Use Schaum s to shorten your study time and get your best test scores      Monograph Series ,1928      Monthly Record of Scientific Literature ,1897

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