

Section 12.2

Chemical Calculations

OBJECTIVES:

- *Construct* "mole ratios" from balanced chemical equations, and *apply* these ratios in mole-mole stoichiometric calculations.
- *Calculate* stoichiometric quantities from balanced chemical equations using units of moles, mass, representative particles, and volumes of gases at STP.

Mole to Mole conversions



- each time we use 2 moles of Al_2O_3 we will also make 3 moles of O_2

These are the two possible conversion factors to use in the solution of the problem.

Mole to Mole conversions

- How many moles of O_2 are produced when 3.34 moles of Al_2O_3 decompose?



If you know the amount of **ANY** chemical in the reaction, you can find the amount of **ALL** the other chemicals!

Practice:



- If 3.84 moles of C_2H_2 are burned, how many moles of O_2 are needed?

- How many moles of C_2H_2 are needed to produce 8.95 mole of H_2O ?

Chemistry Chapter 12 Stoichiometry Notes

Georgia Institute of Technology



Chemistry Chapter 12 Stoichiometry Notes:

Reviewing **Chemistry Chapter 12 Stoichiometry Notes**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is actually astonishing. Within the pages of "**Chemistry Chapter 12 Stoichiometry Notes**," an enthralling opus penned by a very acclaimed wordsmith, readers embark on an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve into the book's central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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