

- silver nitrate + sodium chloride  $\longrightarrow$  silver chloride + sodium nitrate



The ionic equation is:  $\text{Ag}^+_{(aq)} + \text{Cl}^-_{(aq)} \longrightarrow \text{AgCl}_{(s)}$

The spectator ions are  $\text{NO}_3^-$  and  $\text{Na}^+$

(b) Lead (II) iodide can be made by mixing lead (II) nitrate solution with potassium iodide solution.

- Lead (II) nitrate + potassium iodide  $\longrightarrow$  lead (II) iodide + potassium nitrate



- $\text{Pb}(\text{NO}_3)_2_{(aq)} + 2\text{KI}_{(aq)} \longrightarrow \text{PbI}_{2(s)} + 2\text{KNO}_3_{(aq)}$
- The ionic equation is:  $\text{Pb}^{2+}_{(aq)} + 2\text{I}^-_{(aq)} \longrightarrow \text{PbI}_{2(s)}$
- The spectator ions are  $\text{NO}_3^-$  and  $\text{K}^+$

(c) Calcium carbonate forms on e.g. mixing calcium chloride and sodium carbonate solutions

- Calcium chloride + sodium carbonate  $\longrightarrow$  calcium carbonate + sodium chloride



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**Reginald Thomas Booth GLASSPOOL**

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