

Simple Machines

Mechanical Advantage (MA)

$$IMA = \frac{D_E}{D_R} \qquad AMA = \frac{F_R}{F_E}$$

$$\% \text{ Efficiency} = \left(\frac{AMA}{IMA} \right) 100$$

IMA = Ideal Mechanical Advantage

AMA = Actual Mechanical Advantage

D_E = Effort Distance D_R = Resistance Distance

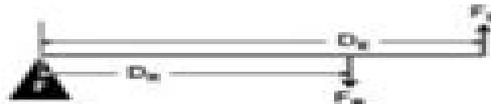
F_E = Effort Force F_R = Resistance Force

Lever

1st Class



2nd Class



3rd Class



Wheel and Axle

Effort at Axle



Effort at Wheel



Pulley Systems

IMA = Total number of strands of a single string supporting the resistance

$$IMA = \frac{D_E \text{ (string pulled)}}{D_R \text{ (resistance lifted)}}$$

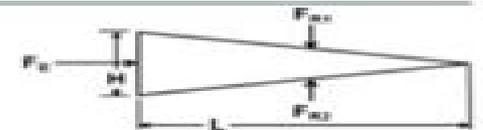
Inclined Plane

$$IMA = \frac{L \text{ (slope)}}{H}$$



Wedge

$$IMA = \frac{L \text{ (L to height)}}{H}$$



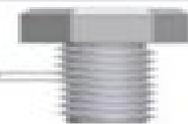
Screw

$$IMA = \frac{C}{\pi r H}$$

$$\text{Pitch} = \frac{1}{\text{TPI}}$$



Pitch



C = Circumference

r = radius

Pitch = distance between threads

TPI = Threads Per Inch

Compound Machines

$$MA_{\text{TOTAL}} = (MA_1) (MA_2) (MA_3) \dots$$

Gears; Sprockets with Chains; and Pulleys with Belts Ratios

$$GR = \frac{N_{\text{out}}}{N_{\text{in}}} = \frac{d_{\text{out}}}{d_{\text{in}}} = \frac{\omega_{\text{in}}}{\omega_{\text{out}}} = \frac{T_{\text{out}}}{T_{\text{in}}}$$

$$\frac{d_{\text{out}}}{d_{\text{in}}} = \frac{\omega_{\text{in}}}{\omega_{\text{out}}} = \frac{T_{\text{out}}}{T_{\text{in}}} \text{ (pulleys)}$$

Compound Gears

$$GR_{\text{TOTAL}} = \left(\frac{B}{A} \right) \left(\frac{D}{C} \right)$$



GR = Gear Ratio

ω_{in} = Angular Velocity - driver

ω_{out} = Angular Velocity - driven

N_{in} = Number of Teeth - driver

N_{out} = Number of Teeth - driven

d_{in} = Diameter - driver

d_{out} = Diameter - driven

T_{in} = Torque - driver

T_{out} = Torque - driven

Civil Engineering Material Quantity Formulas

**Xiu Li Du, Jian Jun Zheng, Wei Ming
Yan, Yue Li, Jian Wei Zhang**

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web ph titration graph 1 lab data provide columns a and b 2 a graph of ph column b plotted as ordinate vs volume column a as abscissa provides a titration curve as illustrated by graph 1 or 2 below 3 estimate of the midpoint of the steeply rising section of the titration curve provides an evaluation of the equivalence point

acid base titration lab report 260678496 pdf studocu - Feb 16 2022

web acid base titration lab report jessica chung chem 267 lab date completed 25 september 2019 date turned in 9 october 2019 figure 1 acid base titration curve from titration with indicator for part 1 volume of naoh added ml ph of the solution 4 5 4 6 4 6 5 9 5 9

experiment 3 acid base titration chem 1202 expt 3 title - Sep 25 2022

web acid base reactions can be used to determine concentrations of an acid or a base using stoichiometric calculations and titrimetric analysis the volume of a solution required to react

acid base titration curves 1 1 microlab inc - Apr 01 2023

web background plot of the ph of a solution against the volume of titrant added is called a titration curve the ph can be measured directly with a ph meter while titrant is added from a buret for the acids used in this experiment the titrant will always be a 0.1 M solution of the strong base NaOH

17 3 acid base titrations chemistry libretexts - Oct 27 2022

web oct 30 2022 there are two basic types of acid base titrations indicator and potentiometric in an indicator based titration you add another chemical that changes color at the ph equal to the equivalence point when the acid and base are in
titration screen experiment resource rsc education - Aug 25 2022

web rp 1 make up a volumetric solution and carry out a simple acid base titration physical chemistry acids and bases ph curves titrations and indicators titrations of acids with bases students should be able to perform calculations for these titrations based on experimental results inorganic chemistry transition metals variable

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Weak strong acids and strong bases completely ionize in solution resulting in water and a salt for a weak acid there is only partial ionization the titration curve of a strong base strong acid showed a stretched out curve as it started with a slow gradual change in pH as it reached the equivalence point