

KINETICS Practice Problems and Solutions

Name: _____
 Period: _____ Class: _____

AP Chemistry
 Dr. McBride

The following questions represent general types of rate questions. Please answer each question completely and thoroughly. The solutions will be posted on the my Moodle page.

1. Consider the reaction: $2\text{H}_2\text{O} \rightarrow 2\text{H}_2 + \text{O}_2$. A rate study of this reaction was conducted at 298 K. The data that were obtained are shown in the table.

$[\text{H}_2\text{O}]$, mol/L	$[\text{O}_2]$, mol/L	Initial Rate, mol/L·s
0.0100	0.0000	1.20×10^{-2}
0.0200	0.0000	4.80×10^{-2}
0.0300	0.0000	8.10×10^{-2}

- a. Write the rate law for the reaction. $\text{Rate} = k[\text{H}_2\text{O}]^2$
 $k = 1.20 \times 10^{-2} \text{ s}^{-1}$
- b. Write the rate law for the reaction. $\text{Rate} = k[\text{H}_2\text{O}]^2$
- c. Determine the order and units of the rate constant. $k = 1.20 \times 10^{-2} \text{ s}^{-1}$ (order 2, units s^{-1})

2. Consider the reaction: $2\text{H}_2\text{O} \rightarrow 2\text{H}_2 + \text{O}_2$. A rate study of this reaction was conducted at 298 K. The data that were obtained are shown in the table.

$[\text{H}_2\text{O}]$, mol/L	$[\text{O}_2]$, mol/L	Initial Rate, mol/L·s
0.010	0.000	0.010
0.020	0.000	0.040
0.030	0.000	0.090

- a. Write the rate law for the reaction. $\text{Rate} = k[\text{H}_2\text{O}]^2$
 $k = 1.00 \text{ s}^{-1}$
- b. Write the rate law for the reaction. $\text{Rate} = k[\text{H}_2\text{O}]^2$
- c. Determine the order and units of the rate constant. $k = 1.00 \text{ s}^{-1}$ (order 2, units s^{-1})

3. Consider the following reactions: $\text{A}_2 + \text{B}_2 \rightarrow \text{A}_2\text{B} + \text{B}$ (slow)
 $\text{A}_2 + \text{B} \rightarrow \text{A} + \text{A} + \text{B}$ (fast)

- a. Write the overall balanced chemical equation: $2\text{A}_2 + \text{B}_2 \rightarrow 2\text{A} + \text{A}_2\text{B}$
- b. Identify any intermediates within the mechanism: B
- c. What is the order with respect to each reactant? $\text{A}_2: 2; \text{B}_2: 1$

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RM Cervero



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