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$$m = \frac{M \cdot 1000}{1000d}$$

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solution of molarity 1.5 M requires 1.5 mol of Na to every litre of solvent

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$$M_{mix} = \frac{m_1 \cdot V_1 + m_2 \cdot V_2 + \dots + m_n \cdot V_n}{V_{total}}$$

1.5 x 480 + 1.2 x 520 = 480 + 520 = 1000 ml
1.344 M with increase in temperature which of these changes a molality b fraction of solute present in water

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