

Name _____ Date _____ Period _____

Acid Bases PHET WebLab for Chromatography

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<https://phet.colorado.edu/en/simulations/versions/acid-base-solutions> Click <Introduction> to begin.

Part 2: Procedure

1. The lab has 2 tools that allow you to test for pH values: A probe  and pH paper . Use each one by dipping it into the solution to be tested. Try all the given types of solutions and fill in the Data Chart with the pH value 0-14.

2. The circuit with a battery and bulb as shown:  is the tool used to test for conduction of a solution. By dipping the wire leads into the solution, the bulb will either remain unlit, be dimly lit, be somewhat bright or very bright. Test each solution and record your observation for the bulb's brightness in the chart below.

Part 1: Data	pH Value from Probe	Color & pH Value from pH Paper	Observations from Circuit Tool Describe the brightness
Water (H ₂ O) 			
Strong Acid (HCl) 			
Weak Acid (H ₂ CO ₃) 			
Strong Base (NaOH) 			
Weak Base (NH ₃) 			

Part 2: Analysis

- What pH value range is observed: a. for acids? _____ b. for bases? _____
- Why are some solutions better conductors of electricity?

Part 2 Procedure, Data & Analysis:

Recall: The amount of ionization or dissociation of ions determines the strength of an acid or base. The concentration of [H₃O⁺], hydronium ion and [OH⁻], hydroxide ion, can be used to calculate pH and pOH as shown on the diagram here.

Note: we use [H₃O⁺] and [H⁺] interchangeably.

- Click on Water Solution, Graph View, Probe Tool. Insert the probe in the water. Notice that the initial



Acid And Bases Ph Phet Lab Answers

Michael Seilmaier



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