

# 11-3 Practice Problems

limiting reactants - part 2

Name: \_\_\_\_\_

Date: \_\_\_\_\_

1. If 4.1 g of Cr is heated with 9.3 g of  $\text{Cl}_2$ , what mass of  $\text{CrCl}_3$  will be produced?



$$4.1\text{g} \times \frac{1\text{mole}}{52\text{g}} = 0.078\text{moles Cr}$$

$$9.3\text{g} \times \frac{1\text{mole}}{70\text{g}} = 0.133\text{moles Cl}_2$$

$$0.078\text{moles} \times \frac{3}{2} = 0.1182\text{moles Cl}_2 \text{ needed. Do I have it?}$$

2. What mass of  $\text{SO}_2$  is produced from the reaction between 31.5 g of  $\text{S}_8$  and 8.65 g of  $\text{O}_2$ ?



3. What mass of  $\text{SO}_2$  is produced from the reaction of 12.4 g of  $\text{SO}_2$  and 3.45 g of  $\text{O}_2$ ?



4. What mass of  $\text{SO}_2$  is produced from the reaction of 58 g of  $\text{SO}_2$  and 1.64 g of  $\text{H}_2\text{O}$ ?



## 11 3 Practice Answers

**Frederic Butterfield Knight, John Ward  
Studebaker, Giles Murrel Ruch**

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