

The Activity Series

Use with Chapter 11
Section 11.2

1. The ions of the following pairs of elements, including the ions that could replace the other in aqueous solution:

- | | |
|-----------------|---------------|
| a. calcium ion | e. zinc ion |
| b. sodium ion | f. sodium ion |
| c. aluminum ion | g. silver ion |
| d. iron ion | |

2. The ions of the following reactions are the products when the reactants, including the cations, react completely in an aqueous solution. Write the products of the reaction, if any, using the appropriate symbols for the products of the reaction. If no reaction occurs, place the hydrogen ion from the reactants on the product side.

- | | |
|----------------------------|----------------|
| a. Lead + Bromine | → $PbBr_2$ |
| b. Lead + Nitric acid | → $Pb(NO_3)_2$ |
| c. Silver + Nitric acid | → $AgNO_3$ |
| d. Zinc + Nitric acid | → $Zn(NO_3)_2$ |
| e. Magnesium + Nitric acid | → $Mg(NO_3)_2$ |
| f. Iron + Nitric acid | → $Fe(NO_3)_3$ |

3. Magnesium metal reacts with various acids. Write the balanced chemical equation for the reaction between magnesium metal and each of the acids listed below. Write the products of the reaction, if any, using the appropriate symbols for the products of the reaction.

Hydrochloric acid: $Mg + 2HCl \rightarrow MgCl_2 + H_2$

Nitric acid: $Mg + 2HNO_3 \rightarrow Mg(NO_3)_2 + H_2$

4. Write the balanced chemical equation for the reaction between iron metal and each of the acids listed below. Write the products of the reaction, if any, using the appropriate symbols for the products of the reaction.

Hydrochloric acid: $Fe + 2HCl \rightarrow FeCl_2 + H_2$

Nitric acid: $Fe + 4HNO_3 \rightarrow Fe(NO_3)_3 + 2H_2O + NO$

5. Write the balanced chemical equation for the reaction between zinc metal and each of the acids listed below. Write the products of the reaction, if any, using the appropriate symbols for the products of the reaction.

$Zn + 2HCl \rightarrow ZnCl_2 + H_2$

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Camilla Rothe



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