



Chemistry

Name: _____

Section _____

Charles' Law WS

Date: _____

Solve the following problems. Assume constant temperature. Show setups, the formula, all substitutions, and include all units and cancellations.

1. A sample of nitrogen gas occupies a volume of 250. mL at 25°C. What volume will it occupy at 95°C?

$$V_1 = 250. \text{ mL} \quad T_1 = 25^\circ\text{C} = 298 \text{ K} \quad T_2 = 95^\circ\text{C} = 368 \text{ K} \quad V_2 = ?$$

$$\frac{V_1}{T_1} = \frac{V_2}{T_2} \quad V_2 = \frac{V_1 T_2}{T_1} \quad V_2 = \frac{250. \text{ mL} \times 368 \text{ K}}{298 \text{ K}} = 309 \text{ mL}$$

2. Oxygen gas is at a temperature of 40°C when it occupies a volume of 2.3 liters. To what temperature should it be raised to occupy a volume of 6.5 liters?

$$V_1 = 2.3 \text{ L} \quad T_1 = 40^\circ\text{C} = 313 \text{ K} \quad V_2 = 6.5 \text{ L} \quad T_2 = ?$$

$$\frac{V_1}{T_1} = \frac{V_2}{T_2} \quad T_2 = \frac{V_2 T_1}{V_1} \quad T_2 = \frac{6.5 \text{ L} \times 313 \text{ K}}{2.3 \text{ L}} = 880 \text{ K} = 610^\circ\text{C}$$

3. Hydrogen gas was cooled from 151°C to 51°C. Its new volume is 75 mL. What was the original volume?

$$T_1 = 151^\circ\text{C} = 424 \text{ K} \quad T_2 = 51^\circ\text{C} = 324 \text{ K} \quad V_2 = 75 \text{ mL} \quad V_1 = ?$$

$$\frac{V_1}{T_1} = \frac{V_2}{T_2} \quad V_1 = \frac{V_2 T_1}{T_2} \quad V_1 = \frac{75 \text{ mL} \times 424 \text{ K}}{324 \text{ K}} = 98 \text{ mL}$$

4. Chlorine gas occupies a volume of 25 mL at 305 K. What volume will it occupy at 605 K?

$$V_1 = 25 \text{ mL} \quad T_1 = 305 \text{ K} \quad T_2 = 605 \text{ K} \quad V_2 = ?$$

$$\frac{V_1}{T_1} = \frac{V_2}{T_2} \quad V_2 = \frac{V_1 T_2}{T_1} \quad V_2 = \frac{25 \text{ mL} \times 605 \text{ K}}{305 \text{ K}} = 50. \text{ mL}$$

5. A sample of neon gas at 55°C and a volume of 2.5 liters is cooled to 25°C. What is the new volume?

$$T_1 = 55^\circ\text{C} = 328 \text{ K} \quad V_1 = 2.5 \text{ L} \quad T_2 = 25^\circ\text{C} = 298 \text{ K} \quad V_2 = ?$$

$$\frac{V_1}{T_1} = \frac{V_2}{T_2} \quad V_2 = \frac{V_1 T_2}{T_1} \quad V_2 = \frac{2.5 \text{ L} \times 298 \text{ K}}{328 \text{ K}} = 2.3 \text{ L}$$

6. Fluorine gas at 305 K occupies a volume of 525 mL. What temperature would result in a volume of 310 mL?

$$T_1 = 305 \text{ K} \quad V_1 = 525 \text{ mL} \quad V_2 = 310 \text{ mL} \quad T_2 = ?$$

$$\frac{V_1}{T_1} = \frac{V_2}{T_2} \quad T_2 = \frac{V_2 T_1}{V_1} \quad T_2 = \frac{310 \text{ mL} \times 305 \text{ K}}{525 \text{ mL}} = 180 \text{ K}$$

7. Helium occupies a volume of 3.8 liters at -45°C. What volume will it occupy at 45°C?

$$V_1 = 3.8 \text{ L} \quad T_1 = -45^\circ\text{C} = 228 \text{ K} \quad T_2 = 45^\circ\text{C} = 318 \text{ K} \quad V_2 = ?$$

$$\frac{V_1}{T_1} = \frac{V_2}{T_2} \quad V_2 = \frac{V_1 T_2}{T_1} \quad V_2 = \frac{3.8 \text{ L} \times 318 \text{ K}}{228 \text{ K}} = 5.3 \text{ L}$$

8. A sample of argon gas is cooled and its volume changes from 385 mL to 255 mL. If its final temperature is -55°C, what was its original temperature?

$$V_1 = 385 \text{ mL} \quad V_2 = 255 \text{ mL} \quad T_2 = -55^\circ\text{C} = 218 \text{ K} \quad T_1 = ?$$

$$\frac{V_1}{T_1} = \frac{V_2}{T_2} \quad T_1 = \frac{V_1 T_2}{V_2} \quad T_1 = \frac{385 \text{ mL} \times 218 \text{ K}}{255 \text{ mL}} = 329 \text{ K} = 56^\circ\text{C}$$

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Clifford Lane



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