

Name: \_\_\_\_\_

**MULTIPLE CHOICE.** Choose the one alternative that best completes the statement or answers the question.

- 1) Ham radio operators often broadcast on the 6-meter band. The frequency of this electromagnetic radiation is \_\_\_\_\_ MHz.  
A) 50                      B) 20                      C) 2.0                      D) 200                      E) 500
- 2) What is the frequency of light ( $s^{-1}$ ) that has a wavelength of  $1.23 \times 10^{-6}$  cm?  
A)  $2.44 \times 10^{16}$               B)  $1.04 \times 10^{13}$               C)  $9.62 \times 10^{12}$               D) 3.69                      E)  $4.10 \times 10^{-17}$
- 3) The wavelength of a photon that has an energy of  $5.25 \times 10^{-19}$  J is \_\_\_\_\_ m.  
A)  $4.21 \times 10^{-24}$               B)  $2.38 \times 10^{23}$               C)  $2.64 \times 10^6$               D)  $3.79 \times 10^7$               E)  $3.79 \times 10^{-7}$
- 4) The frequency of a photon that has an energy of  $3.7 \times 10^{-18}$  J is \_\_\_\_\_  $s^{-1}$ .  
A)  $5.4 \times 10^{-8}$               B)  $5.6 \times 10^{15}$               C)  $1.8 \times 10^{-16}$               D)  $2.5 \times 10^{-15}$               E)  $2.5 \times 10^{15}$
- 5) A mole of red photons of wavelength 725 nm has \_\_\_\_\_ kJ of energy.  
A)  $6.05 \times 10^{-3}$               B)  $2.74 \times 10^{-19}$               C)  $4.56 \times 10^{-46}$               D) 227                      E) 165
- 6) Of the following transitions in the Bohr hydrogen atom, the \_\_\_\_\_ transition results in the emission of the highest-energy photon.  
A)  $n = 6 \rightarrow n = 1$               B)  $n = 1 \rightarrow n = 6$               C)  $n = 3 \rightarrow n = 6$               D)  $n = 6 \rightarrow n = 3$               E)  $n = 1 \rightarrow n = 4$
- 7) Using Bohr's equation for the energy levels of the electron in the hydrogen atom, determine the energy (J) of an electron in the  $n = 4$  level.  
A)  $-5.45 \times 10^{-19}$               B)  $-1.84 \times 10^{-29}$               C)  $-1.36 \times 10^{-19}$               D)  $+1.84 \times 10^{-29}$               E)  $-7.34 \times 10^{18}$
- 8) The lines in the emission spectrum of hydrogen result from \_\_\_\_\_.  
A) electrons given off by hydrogen as it cools  
B) decomposing hydrogen atoms  
C) electrons given off by hydrogen when it burns  
D) energy given off in the form of visible light when an electron moves from a higher energy state to a lower energy state  
E) protons given off when hydrogen burns
- 9) When the electron in a hydrogen atom moves from  $n = 6$  to  $n = 2$ , light with a wavelength of \_\_\_\_\_ nm is emitted.  
A) 657                      B) 93.8                      C) 411                      D) 434                      E) 487
- 10) What is the De Broglie wavelength (m) of a 2.0 kg object moving at a speed of 50 m/s?  
A)  $1.5 \times 10^{35}$               B)  $3.8 \times 10^{34}$               C)  $2.6 \times 10^{-35}$               D)  $5.3 \times 10^{-33}$               E)  $6.6 \times 10^{-36}$

# A P Chemistry Practice Test Ch 7 Atomic Structure And

**Liying Dong**



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