

## VOLUMETRIC ANALYSIS

### AIM

Determine the concentration of aqueous hydrogen peroxide and the percentage purity of oxalate in an unknown oxalate sample by titration with potassium permanganate.

### INTRODUCTION

Volumetric analysis is a quantitative analysis process whereby volumes are measured. A volume of reagent known as the standard solution of known concentration is chemically reacted with a solution of unknown concentration in order to determine the concentration of the unknown. The solution that is used as standard is called the titrant which is delivered from a burette and is added until the reaction is completed. The end point of the titration is noted by observing the change of colour of an indicator. In a direct titration, the titrant is added to analyte until the end point is observed whereas in a back titration, a known excess of a standard reagent is added to the analyte. A second standard reagent is then used to titrate the excess of the first reagent.

All the calculations in determining the molarity and average concentration for all equations involve the formulae:

$$M_A V_A = a$$

$$M_B V_B = b$$

The equation for the standardization experiment is given in the ionic form as:



The back titration is explained by the equations:



### APPARATUS AND REAGENTS

1.6g of Potassium permanganate

Boiling distilled water

Distilled water

Erlenmeyer flask

Rubber stopper

Sodium oxalate

Sulphuric acid

Source of heat

Burette

Hydrogen peroxide

# A Volumetric Analysis Lab Report Answers

**Johns Hopkins University**



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