

Use your notes from the Atomic Structure program to answer the following questions.

- The atomic number tells the number of positively charged protons in the nucleus of an atom. The atom is neutral because this is also the number of negatively charged electrons in the atom.
- The mass number tells the total number of protons and neutrons in the nucleus of an atom. These particles collectively are called nucleons since both are located in the nucleus.
- Isotopes are atoms of the same element with different numbers of neutrons which results in different mass numbers.
- Write the hyphen notation of the three isotopes of hydrogen:
 - Hydrogen - 1 ${}^1_1\text{H}$
 - Hydrogen - 2 ${}^2_1\text{H}$
 - Hydrogen - 3 ${}^3_1\text{H}$
- Fill in the following chart (Hint: Isotopes may have different masses than the periodic table lists!):

Symbol	Atomic #	Mass #	# of protons	# of electrons	# of neutrons
Hg	80	201	80	80	121
Mo	42	96	42	42	54
Cu	29	64	29	29	35
C	6	14	6	6	8
	15	31	15	15	16
Pb	82	208	82	82	126
Na	11	24	11	11	13

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Table of Contents Chapter 4 Atomic Structure Worksheet Answers Pearson

1. Understanding the eBook Chapter 4 Atomic Structure Worksheet Answers Pearson
 - The Rise of Digital Reading Chapter 4 Atomic Structure Worksheet Answers Pearson
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 4 Atomic Structure Worksheet Answers Pearson
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 4 Atomic Structure Worksheet Answers Pearson
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 4 Atomic Structure Worksheet Answers Pearson
 - Personalized Recommendations
 - Chapter 4 Atomic Structure Worksheet Answers Pearson User Reviews and Ratings
 - Chapter 4 Atomic Structure Worksheet Answers Pearson and Bestseller Lists
5. Accessing Chapter 4 Atomic Structure Worksheet Answers Pearson Free and Paid eBooks
 - Chapter 4 Atomic Structure Worksheet Answers Pearson Public Domain eBooks
 - Chapter 4 Atomic Structure Worksheet Answers Pearson eBook Subscription Services
 - Chapter 4 Atomic Structure Worksheet Answers Pearson Budget-Friendly Options

6. Navigating Chapter 4 Atomic Structure Worksheet Answers Pearson eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 4 Atomic Structure Worksheet Answers Pearson Compatibility with Devices
 - Chapter 4 Atomic Structure Worksheet Answers Pearson Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 4 Atomic Structure Worksheet Answers Pearson
 - Highlighting and Note-Taking Chapter 4 Atomic Structure Worksheet Answers Pearson
 - Interactive Elements Chapter 4 Atomic Structure Worksheet Answers Pearson
8. Staying Engaged with Chapter 4 Atomic Structure Worksheet Answers Pearson
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 4 Atomic Structure Worksheet Answers Pearson
9. Balancing eBooks and Physical Books Chapter 4 Atomic Structure Worksheet Answers Pearson
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 4 Atomic Structure Worksheet Answers Pearson
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter 4 Atomic Structure Worksheet Answers Pearson
 - Setting Reading Goals Chapter 4 Atomic Structure Worksheet Answers Pearson
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 4 Atomic Structure Worksheet Answers Pearson
 - Fact-Checking eBook Content of Chapter 4 Atomic Structure Worksheet Answers Pearson
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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