

## IONIC BONDING

Ionic bonding occurs when a metal transfers one or more electrons to a nonmetal in an effort to obtain a stable octet of electrons. For example, the transfer of an electron from sodium to chlorine can be depicted by a Lewis dot diagram.



Calcium would need two chlorine atoms to get rid of its two valence electrons.



Show the transfer of electrons in the following combinations.

1. K + F



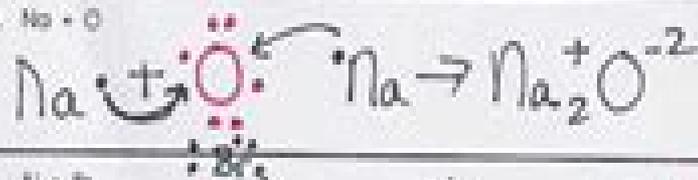
2. Mg + I



3. Be + S



4. Na + O



5. Al + Br



## COVALENT BONDING

Name \_\_\_\_\_

Covalent bonding occurs when two or more nonmetals share electrons, attempting to gain a stable octet of electrons at least part of the time. For example:

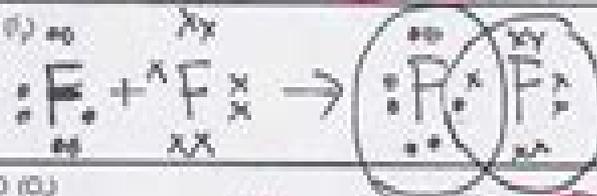


Show how covalent bonding occurs in each of the following pairs of atoms. Atoms may share one, two or three pairs of electrons.

1. H + H (H<sub>2</sub>)



2. F + F (F<sub>2</sub>)



3. O + O (O<sub>2</sub>)



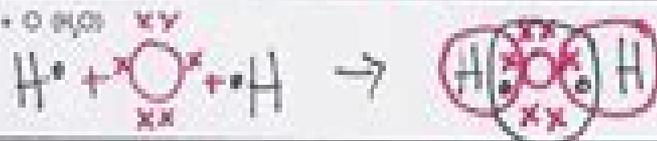
4. N + N (N<sub>2</sub>)



5. C + O (CO<sub>2</sub>)



6. H + O (H<sub>2</sub>O)



# Chapter 8 Covalent Bonding And Molecular Structure

**Paul B. Kelter**



## Chapter 8 Covalent Bonding And Molecular Structure:

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