

CHAPTER 6 REVIEW*Chemical Bonding***MIXED REVIEW**

SHORT ANSWER Answer the following questions in the space provided.

1. Name the type of energy that is a measure of strength for each of the following types of bonds:

| | | | |
|-------|---------------------------------|-------|------------------|
| _____ | lattice energy | _____ | a. ionic bond |
| _____ | bond energy | _____ | b. covalent bond |
| _____ | enthalpy of vaporization | _____ | c. metallic bond |

2. Use the electronegativity values shown in **Figure 20**, on page 161 of the text, to determine whether each of the following bonds is nonpolar covalent, polar covalent, or ionic.

| | | | | | |
|-------|-----------------------|----------|-------|--------------------------|--------|
| _____ | ionic | a. H—F | _____ | nonpolar covalent | d. H—H |
| _____ | ionic | b. Na—Cl | _____ | polar covalent | e. H—C |
| _____ | polar covalent | c. H—O | _____ | polar covalent | f. H—N |

3. How is a hydrogen bond different from an ionic or covalent bond?

A hydrogen bond is a dipole-dipole attraction between a partially positive hydrogen atom and the unshared electron pair of a strongly electronegative atom such as O, N, or F. Unlike ionic or covalent bonds, in which electrons are given up or shared, the hydrogen bond is a weaker attraction. Hydrogen bonds are generally intermolecular, while ionic and covalent bonds occur between ions or atoms respectively.

4. H₂S and H₂O have similar structures and their central atoms belong to the same group. Yet H₂S is a gas at room temperature and H₂O is a liquid. Use bonding principles to explain why this is.

Oxygen has higher electronegativity than sulfur, which creates a highly polar bond. Increased polarity in H₂O bonds means a stronger intermolecular attraction, making water a liquid at room temperature. Hydrogen bonding exists between water molecules, but not between hydrogen sulfide molecules.

Chapter 8 Covalent Bonding Answers Key

John H. Meiser



Chapter 8 Covalent Bonding Answers Key:

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