

10

CHEMICAL QUANTITIES

Practice Problems

On a piece of paper, solve the following problems.

YOU MUST SHOW YOUR WORK TO GET CREDIT FOR YOUR ANSWERS!!!

SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER

1. What is the molar mass of sucrose ($C_{12}H_{22}O_{11}$)?
2. What is the molar mass of each of the following compounds?
 - a. phosphorus pentachloride (PCl_5)
 - b. uranium hexafluoride (UF_6)
3. Calculate the molar mass of each of the following ionic compounds:
 - a. $KMnO_4$
 - b. $Ca_3(PO_4)_2$
4. How many moles is 3.52×10^{24} molecules of water?
5. How many atoms of zinc are in 0.60 mol of zinc?
6. What is the mass of 1.00 mol of oxygen (O_2)?

SECTION 10.2 MOLE–MASS AND MOLE–VOLUME RELATIONSHIPS

1. What is the molar mass of each of the following compounds?
 - a. $C_6H_{12}O_6$
 - b. $NaHCO_3$
 - c. C_2H_{12}
 - d. KNH_4SO_4
2. Calculate the mass in grams of each of the following:
 - a. 8.0 mol lead oxide (PbO)
 - b. 0.75 mol hydrogen sulfide (H_2S)
 - c. 0.00100 mol silicon tetrahydride (SiH_4)
 - d. 1.50×10^{-2} mol molecular oxygen (O_2)
 - e. 2.30 mol ethylene glycol ($C_2H_6O_2$)
3. How many grams are in 1.73 mol of dinitrogen pentoxide (N_2O_5)?
4. How many grams are in 0.658 mol of calcium phosphate [$Ca_3(PO_4)_2$]?
5. Calculate the number of moles in each of the following:
 - a. 0.50 g sodium bromide ($NaBr$)
 - b. 13.5 g magnesium nitrate [$Mg(NO_3)_2$]
 - c. 1.02 g magnesium chloride ($MgCl_2$)
 - d. 0.00100 g monochloromethane (CH_3Cl)
 - e. 1.50×10^{-3} g propylene glycol [$C_3H_8(OH)_2$]
6. A chemist plans to use 435.0 grams of ammonium nitrate (NH_4NO_3) in a reaction. How many moles of the compound is this?
7. A solution is to be prepared in a laboratory. The solution requires 0.0465 mol of quinine ($C_{20}H_{24}N_2O_2$). What mass, in grams, should the laboratory technician obtain in order to make the solution?
8. What is the volume at STP of 2.66 mol of methane (CH_4) gas?
9. How many moles is 135 L of ammonia (NH_3) gas at STP?

Chemistry Chapter 10 The Mole Study Guide Answers

James E. Brady, John R. Holum



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bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

Know Your 'O' Level Chemistry - A Study Guide, **Study Guide with Answers to Selected Problems** H. Stephen Stoker, Danny V. White, 2003-04 This supplement includes for each chapter a brief overview activities and practice problems to reinforce skills and a practice test The answers section includes answers for all odd numbered end of chapter exercises

Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH_4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

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Study Guide and Solutions Manual, Fundamentals of General, Organic, and Biological Chemistry, Third Edition John McMurry, Susan McMurry, 1999 Provides worked out solutions to text problems along with chapter by chapter outlines and a variety of self tests at the end of each chapter

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