

## Section 12.2

### Chemical Calculations

#### **OBJECTIVES:**

- *Construct* "mole ratios" from balanced chemical equations, and *apply* these ratios in mole-mole stoichiometric calculations.
- *Calculate* stoichiometric quantities from balanced chemical equations using units of moles, mass, representative particles, and volumes of gases at STP.

#### **Mole to Mole conversions**



- each time we use 2 moles of  $\text{Al}_2\text{O}_3$  we will also make 3 moles of  $\text{O}_2$

These are the two possible conversion factors to use in the solution of the problem.

#### **Mole to Mole conversions**

- How many moles of  $\text{O}_2$  are produced when 3.34 moles of  $\text{Al}_2\text{O}_3$  decompose?



If you know the amount of **ANY** chemical in the reaction, you can find the amount of **ALL** the other chemicals!

#### **Practice:**



- If 3.84 moles of  $\text{C}_2\text{H}_2$  are burned, how many moles of  $\text{O}_2$  are needed?
  
- How many moles of  $\text{C}_2\text{H}_2$  are needed to produce 8.95 mole of  $\text{H}_2\text{O}$ ?

# Chemistry Chapter 12 Stoichiometry Notes

**Puja Mehta**

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