

Chapter 13 - Practical Work

(a) (i) Chemical test for ammonia:

If a rod dipped in concentrated hydrochloric acid is brought near ammonia gas, dense white fumes of ammonium chloride (NH_4Cl) are formed.

(ii) Chemical test for Sulphur dioxide:

It decolorizes pink coloured potassium permanganate solution.



(iii) Chemical test for HCl:

When HCl gas is passed through AgNO_3 solution, white precipitates of AgCl are formed which gets dissolved in excess of NH_4OH .



(iv) Chemical test for Chlorine:

It turns moist starch iodide paper (KI + starch solution) blue black.

(v) Chemical test for Carbon dioxide:

When this gas is passed through lime water, it turns milky due to the formation of white precipitates of CaCO_3 and on passing excess of carbon dioxide gas, this milkiness disappears.

(vi) Chemical test for oxygen:

This gas is absorbed in colourless alkaline solution of pyrogallol and turns it dark brown.

(vii) Chemical test for hydrogen:

It burns with a pop sound when a burning taper is brought near it.

(b) Ammonia is a basic gas and its basic nature is suspected through litmus paper test because it changes the colour of red litmus paper to blue.

(c) Chlorine, carbon dioxide, hydrogen chloride, hydrogen sulphide and sulphur dioxide are acidic gases since they convert blue litmus to red.

(d) A is chlorine and B is Sulphur dioxide.

(e) Water vapour.

Chemistry Chapter 13 Solutions Manual

JS Bruner



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