

1.2 Mole Conversions Solutions

PRACTICE EXAMPLES- To complete the practice examples (1) write the type of question you have, (2) write the equation with units, (3) show the cancellation of units and (4) calculate the solution.

1. Determine how many atoms are present in 2.5 moles of Silicon? (15.05×10^{23} atoms)

MOLES → ATOMS
$$\frac{2.5 \text{ moles}}{1} \times \frac{6.022 \times 10^{23} \text{ atoms}}{1 \text{ mole}} = 15.05 \times 10^{23} \text{ atoms}$$

2. How many moles of Silicon is 3.01×10^{23} atoms? (.5 moles)

ATOMS → MOLES
$$\frac{3.011 \times 10^{23} \text{ atoms}}{1} \times \frac{1 \text{ mole}}{6.022 \times 10^{23} \text{ atoms}} = .5 \text{ moles}$$

3. What is the mass of 4 moles of Iron? (224 grams)

MOLES → GRAMS
$$\frac{4 \text{ moles}}{1} \times \frac{55.9 \text{ grams}}{\text{mole}} = 224 \text{ grams of C}$$

4. How many atoms are in 48.6 grams of Magnesium? (12.04×10^{23} atoms)

GRAMS → ATOMS
$$\frac{48.6 \text{ grams}}{1} \times \frac{1 \text{ mole}}{24.3 \text{ grams}} \times \frac{6.022 \times 10^{23} \text{ atoms}}{1 \text{ mole}} = 12.04 \times 10^{23} \text{ atoms}$$

5. How many grams are in 18.06×10^{23} atoms of Copper? (190.5 grams)

ATOMS → GRAMS
$$\frac{18.06 \times 10^{23} \text{ atoms}}{1} \times \frac{1 \text{ mole}}{6.022 \times 10^{23} \text{ atoms}} \times \frac{63.5 \text{ grams}}{1 \text{ mole}} = 190.5 \text{ grams}$$

6. How many moles of Nickel is 176 grams? (3 moles)

GRAMS → MOLES
$$\frac{176 \text{ grams of C}}{1} \times \frac{\text{mole}}{58.7 \text{ grams}} = 3 \text{ moles}$$

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John McMurry



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Survival Guide to General Chemistry Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics. The focus is not to recap or review the theoretical concepts well described in the available texts. Instead, the topics and descriptions in this book make available specific, detailed, step-by-step methods and procedures for solving the major types of problems in general chemistry. Explanations, instructional process sequences, solved examples, and completely solved practice problems are greatly expanded, containing significantly more detail than can usually be devoted to in a comprehensive text. Many chapters also provide alternative viewpoints as an aid to understanding. Key Features: The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester. Each is written in a specific and detailed step-by-step process for problem solving, whether mathematical or conceptual. Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts. Includes a chapter designed to eliminate confusion concerning acid-base reactions, which often persists through working with acid-base equilibrium. Many chapters provide alternative viewpoints as an aid to understanding. This book addresses a very real need for a large number of incoming freshmen in STEM fields.

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