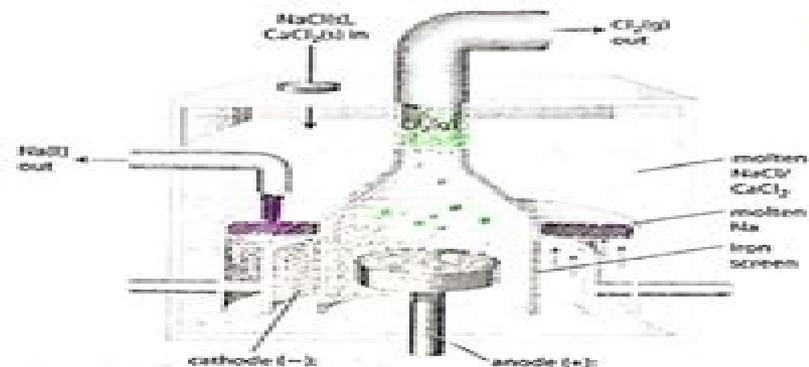


# IGCSE – HONORS – ELECTROCHEMISTRY NOTES AND STUDY GUIDE:

## Electrolysis of sodium chloride (in the industry)

Industrial electrolysis of molten Sodium Chloride requires extreme temperatures.

- Calcium Chloride is added to the mixture to lower the melting point of Sodium Chloride.
- Molten sodium metal floats in the mixture because of its low density.
- The anode is carbon
- The cathode is iron
- Chlorine gas is collected separately



3.

## 4. electrolysis of Aluminum oxide (molten)

Electrodes: graphite (carbon)

Electrolyte: molten Alumina (Al<sub>2</sub>O<sub>3</sub>)

Alumina = Purified Bauxite

Which substance is added to the molten alumina? cryolite

Ions Present: Al<sup>3+</sup>, O<sup>2-</sup>

Half equations:

In the Cathode (Al<sup>3+</sup> + 3e<sup>-</sup> → Al) × 4

In the Anode (2O<sup>2-</sup> → O<sub>2</sub> + 4e<sup>-</sup>) × 3

Overall Equation 4 Al<sup>3+</sup> + 6 O<sup>2-</sup> → 4 Al + 3 O<sub>2</sub>

Observations:

Cathode: The Molten Aluminum can be observed

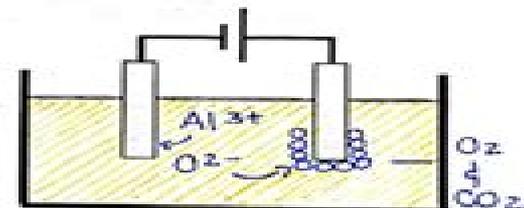
Anode: Bubbles of Oxygen gas will appear

Why is it necessary to change the electrodes periodically?

Because the anodes will burn out. Anode reacts with O<sub>2</sub> to form CO<sub>2</sub>

Which is the purpose of the cryolite?

To lower the melting point of the Aluminum Oxide



# Chemistry Electrochemistry Study Guide Answers Key

**Judd E. Hollander**



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