

Ionic Compounds and Metals

Section 7.1 Ion Formation

pages 206–209

Section 7.1 Assessment

page 209

- 1. Compare** the stability of a lithium atom with that of its ion, Li^+ .

The Li^+ ion is more stable because it has a complete octet.
- 2. Describe** two different causes of the force of attraction in a chemical bond.

The two causes are the attraction between the positive nucleus of one atom and the negative electrons of another atom, and the attraction between positive ions and negative ions.
- 3. Apply** Why are all of the elements from group 18 relatively unreactive, whereas those in group 17 are very reactive?

The group 18 elements, known as noble gases, have complete outer energy levels and do not easily form ions. The group 17 elements are highly reactive because they need only to gain a single electron to form an octet.
- 4. Summarize** ionic bond formation by correctly pairing these terms: *cation*, *anion*, *electron gain*, *electron loss*.

(anion, electron gain), (cation, electron loss)
- 5. Apply** Write out the electron configuration for each atom. Then predict the change that must occur in each to achieve a noble gas configuration.

 - nitrogen
 $[\text{He}]2s^22p^3$; gain 3 electrons (3^- ion) or lose 5 electrons (5^+ ion)
 - sulfur
 $[\text{Ne}]3s^23p^4$; gain 2 electrons (2^- ion)

- barium
 $[\text{Xe}]6s^2$; lose 2 electrons (2^+ ion)
- lithium
 $[\text{He}]2s^1$; lose 1 electron (1^+ ion)

- 6. Model** Draw models to represent the formation of the positive calcium ion and the negative bromide ion.

Models should show that the calcium atom loses two electrons, forming Ca^{2+} , and that bromine gains one electron, forming Br^- . The models should also show the addition of energy to form the calcium ion, and the release of energy to form the bromide ion.

Section 7.2 Ionic Bonds and Ionic Compounds

pages 210–217

Practice Problems

page 212

Explain how an ionic compound forms from these elements.

- sodium and nitrogen
 Three Na atoms each lose $1 e^-$, forming 1^+ ions. One N atom gains $3 e^-$, forming a 3^- ion. The ions attract, forming Na_3N .

$$3 \text{ Na ions } \left(\frac{1+}{\text{Na ion}} \right) + 1 \text{ N ion } \left(\frac{3-}{\text{N ion}} \right)$$

$$= 3(1+) + 1(3-) = 0$$
 The overall charge on one formula unit of Na_3N is zero.
- lithium and oxygen
 Two Li atoms each lose $1 e^-$, forming 1^+ ions. One O atom gains $2 e^-$, forming a 2^- ion. The ions attract, forming Li_2O .

$$2 \text{ Li ions } \left(\frac{1+}{\text{Li ion}} \right) + 1 \text{ O ion } \left(\frac{2-}{\text{O ion}} \right)$$

$$= 2(1+) + 1(2-) = 0$$
 The overall charge on one formula unit of Li_2O is zero.

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Table of Contents Chemistry Matter And Change Chapter 7 Study Guide Answers

1. Understanding the eBook Chemistry Matter And Change Chapter 7 Study Guide Answers
 - The Rise of Digital Reading Chemistry Matter And Change Chapter 7 Study Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Matter And Change Chapter 7 Study Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Matter And Change Chapter 7 Study Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Matter And Change Chapter 7 Study Guide Answers
 - Personalized Recommendations
 - Chemistry Matter And Change Chapter 7 Study Guide Answers User Reviews and Ratings
 - Chemistry Matter And Change Chapter 7 Study Guide Answers and Bestseller Lists
5. Accessing Chemistry Matter And Change Chapter 7 Study Guide Answers Free and Paid eBooks
 - Chemistry Matter And Change Chapter 7 Study Guide Answers Public Domain eBooks
 - Chemistry Matter And Change Chapter 7 Study Guide Answers eBook Subscription Services
 - Chemistry Matter And Change Chapter 7 Study Guide Answers Budget-Friendly Options
6. Navigating Chemistry Matter And Change Chapter 7 Study Guide Answers eBook Formats

- ePub, PDF, MOBI, and More
 - Chemistry Matter And Change Chapter 7 Study Guide Answers Compatibility with Devices
 - Chemistry Matter And Change Chapter 7 Study Guide Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Matter And Change Chapter 7 Study Guide Answers
 - Highlighting and Note-Taking Chemistry Matter And Change Chapter 7 Study Guide Answers
 - Interactive Elements Chemistry Matter And Change Chapter 7 Study Guide Answers
 8. Staying Engaged with Chemistry Matter And Change Chapter 7 Study Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Matter And Change Chapter 7 Study Guide Answers
 9. Balancing eBooks and Physical Books Chemistry Matter And Change Chapter 7 Study Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Matter And Change Chapter 7 Study Guide Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Chemistry Matter And Change Chapter 7 Study Guide Answers
 - Setting Reading Goals Chemistry Matter And Change Chapter 7 Study Guide Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Chemistry Matter And Change Chapter 7 Study Guide Answers
 - Fact-Checking eBook Content of Chemistry Matter And Change Chapter 7 Study Guide Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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