

10. Salts that formed by the reaction of weak acid with strong base are:  
a. Neutral  
b. ✓ Acidic  
c. Basic  
d. Normal
11. Salts that formed by the reaction of weak acid with strong base are:  
a. Acidic  
b. ✓ Basic  
c. Neutral  
d. Normal
12. Alums are:  
a. Single salts  
b. ✓ Double salts  
c. Triple salts  
d. Normal salts
13. The formula of washing soda is.  
a.  $\text{Na}_2\text{CO}_3$   
b.  $\text{Na}_2\text{CO}_3 \cdot 6\text{H}_2\text{O}$   
c. ✓  $\text{Na}_2\text{CO}_3 \cdot 10\text{H}_2\text{O}$   
d.  $\text{NaHCO}_3$
14. All the citrus fruits contain  
a. Acetic acid  
b. ✓ Ascorbic acid  
c. Lactic acid  
d. Butyric acid
15. All the acids and bases react together to form salt and water, the reaction is called.  
a. Hydrolysis  
b. Decomposition  
c. Synthesis  
d. ✓ Neutralization
16. According to Arrhenius, a base gives hydroxide  $\text{OH}^-$  ion in water, while an acid gives  
a. ✓ Hydrogen ion  
b. Metallic ion  
c. Non metallic  
d. Ammonium ion
17. According to Bronsted - Lowry theory, an acid is a proton donor while a base is a/an.  
a. ✓ Proton acceptor  
b. Electron acceptor  
c. Neutron acceptor  
d. None of them
18. According to Lewis, an acid is an acceptor of an electron pair while a base is a donor of:  
a. Proton pair  
b. Neutron pair  
c. ✓ Electron pair  
d. None of them
19. Neutralization is an:  
a. Endothermic reaction  
b. ✓ Exothermic reaction  
c. Addition reaction  
d. Odd reaction
20. Double salts usually have which molecules attached to their crystals:  
a.  $\text{H}_2\text{O}_2$   
b. ✓  $\text{H}_2\text{O}$   
c.  $\text{CO}_2$   
d. CO

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