

# SEMESTER 1 FINAL REVIEW – CHEMISTRY

## 2015-16

Please ensure that you know how to answer these questions correctly, these are all good example problems for your final exam. Study your notes, labs, homework assignments and use the PowerPoint presentations that are online. Any of this information is fair game to be asked on the final exam. Good luck and ask questions in class during the review times if you have them. Don't wait to the last minute to start reviewing!

1. What is the number of significant figures in 1007000.0
2. What is the difference between accuracy and precision?
3. What is Avogadro's number? Why is it useful?
4. Find the molecular weight of  $\text{CO}_2$ ,  $\text{NH}_3$ ,  $\text{SF}_6$  ?
5. Be able to use density to calculate mass and or volume for various objects. Also, understand that density can be used at times with moles and particles through the  $D=m/v$  formula.
6. What is a mole? Be able to explain how the mole is a central connecting idea between mass, particles and gases @STP
7. What is the molar volume of a gas at STP?
8. What are the seven diatomic elements?
9. Define: atom, molecule, diatomic gas, monoatomic gas, element, compound, STP
10. What elements form the following bonds: ionic, covalent
11. Give an example of the following (both name and formula)
  - a. an ionic compound
  - b. a covalent compound
  - c. a metal
  - d. a nonmetal
  - e. a diatomic gas
12. Name the following:

a) NaOH	b) $\text{SO}_2$	c) $\text{C}_2\text{SO}_4$	d) HCl	e) $\text{Ca}(\text{NO}_3)_2$	f) $\text{NH}_4$
g) $\text{N}_2\text{O}_5$	h) $\text{Al}_2(\text{SO}_4)_3$	i) $\text{K}_3\text{PO}_4$	j) $\text{Na}_2\text{O}$	k) CO	l) $\text{CO}_2$
13. Write formulas for following
  - a) dinitrogen trioxide
  - b) potassium sulfate
  - c) hydrogen phosphate
  - d) sodium bromide
  - e) ammonium acetate
  - f) manganese IV nitrate
  - g) iron III hydroxide
  - h) aluminum carbonate
14. Write the following in correct scientific notation:

a) 476	b) 0.0000000002223	c) 3240000000000000
d) $462.9 \times 10^{23}$	e) $0.000000345 \times 10^{12}$	
f) $5677.9 \times 10^{-12}$	g) $0.000000223 \times 10^{-6}$	
15. How many oxygen atoms are there in 420 grams of carbon dioxide?
16. What is the mass of  $1.2 \times 10^{24}$  molecules of bromine?
17. What is the volume occupied by 243 grams of oxygen gas at STP?
18. How do you convert from:
  - a) moles  $\rightarrow$  grams
  - b) grams  $\rightarrow$  moles
  - c) moles  $\rightarrow$  molecules
  - d) moles  $\rightarrow$  liters of gas at STP
  - e) moles  $\rightarrow$  liters of solution

# Chemistry Semester 1 Review Answers

**AW Rasmussen**



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