

Introduction 11

An Electron's Address

Name: _____
 Date: _____
 Class: _____

Introduction: Energy Levels and Sublevels

In addition to the energy level model that you used in the previous lesson, energy levels are subdivided into sublevels. The energy levels are labeled as the letters s, p, d, and f.

Sublevels are labeled with energy levels and the sublevel letter. Each sublevel contains a certain number of orbitals.

Table 1

Energy Level	Number of Sublevels in Each Energy Level
1	1
2	2
3	3
4	4

Note that there is always one s sublevel in each of the 1st, 2nd, 3rd, and 4th energy levels. There are also one p, two d, and three f sublevels in the 2nd, 3rd, and 4th energy levels.

Critical Thinking Questions

- How many sublevels exist in the 1st energy level?
 The sublevel that exists in the 1st energy level is s, and it is the only sublevel.
- How many sublevels exist in the 2nd energy level?
 Two sublevels that exist in the 2nd energy level are s and p sublevels.
- How many sublevels exist in the 3rd energy level?
 Three sublevels that exist in the 3rd energy level are s, p, and d sublevels.
- How many sublevels would you expect to find in the 4th energy level?
 The number of sublevels I would expect to find in the 4th energy level would be 4 sublevels. There would be 1 sublevel s, 3 sublevels p, 5 sublevels d, and 7 sublevels f.
- How many sublevels would you expect to find in the 5th energy level?
 There are 5 sublevels. There are 1 s, 3 p, 5 d, and 7 f sublevels. The 5th energy level would have 1 s, 3 p, 5 d, and 7 f sublevels. There are 1 s, 3 p, 5 d, and 7 f sublevels.

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